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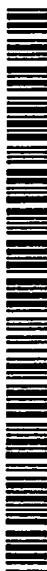
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(54) Title: LIQUID INK USING AN ACID-BASE CROSSLINKED ORGANOSOL

(57) Abstract: Gel organosol dispersions, and high solids color inks based upon these dispersions, featuring a carrier liquid having a Kauri-Butanol number less than 30 and a gel organosol dispersion that includes one or more graft copolymers comprising a (co)polymeric steric stabilizer covalently bonded to a (co)polymeric thermoplastic core that is insoluble in the carrier liquid. The copolymers are crosslinked by the interaction of a Lewis acid functional group and a Lewis base functional group.

## **LIQUID INK USING AN ACID-BASE CROSSLINKED ORGANOSOL**

### **TECHNICAL FIELD**

This invention relates to liquid ink compositions, in particular, to pigments dispersed in gel organosols to provide improved ink compositions and liquid toners for use in ink transfer, ionographic, electrographic and electrophotographic printing processes.

### **BACKGROUND**

Liquid inks are widely used in a variety of printing processes, for example offset, intaglio, rotogravure, ink jet and electrographic printing. Many of the desired characteristics of the pigment dispersions used in the liquid inks are the same for each of the respective processes even though the final ink formulations may be substantially different. For example, the stability of the pigment dispersion both on the shelf and under shear conditions is an important consideration regardless of the final use of the liquid ink. The art continuously searches for more stable pigment dispersions to provide more flexibility in ink formulations, which in turn yields better efficiency and waste reduction in the various printing processes.

Electrographic printing refers to a printing process that uses an applied electric field and charged particles to produce a printed image on a receptor material. The art generally refers to the charged particles as toners. Electrographic printing generally includes electrostatographic printing, ionographic printing, electrophotographic printing and the like. In electrophotographic applications, which include devices such as photocopiers, laser printers, facsimile machines and the like, the toners may be in the form of dry particles or particles dispersed in a carrier liquid. Particles dispersed in a liquid medium for imaging purposes are generally referred to as liquid inks, liquid toners or liquid developers.

Generally, the electrophotographic process includes the steps of forming a latent electrostatic image on a charged photoconductor by exposing the photoconductor to radiation in an imagewise pattern, developing the image by contacting the photoconductor with a liquid developer, and finally transferring the

image to a receptor. The final transfer step may be performed either directly or indirectly through an intermediate transport member. The developed image is usually subjected to heat and/or pressure to permanently fuse the image to the receptor.

In the field of electrographic printing, particularly electrophotographic printing, a variety of both liquid and dry developing compositions have been employed to develop the latent electrostatic images. Dry toner compositions suffer from a number of disadvantages. For example, dry toners are known to be difficult to control during the latent image development and transfer processes; this leads to toner scatter within the printer device and may create excessive amounts of dust and abrasive wear of the printer components. Some dry toner compositions must also be fixed by fusing at elevated temperature, which requires a large source of energy and may limit the choices of receptor materials to which the developed latent image may be transferred. Moreover, dry toners must be triboelectrically charged, which makes the printing process very sensitive to both the temperature and humidity of the ambient air and may result in printing delays due to charge equilibration. The limited particle size of the toner is another disadvantage of dry toners. If the particle size is small, the dry toner can become airborne and create a potential health hazard due to inhalation of the particles. On the other hand, the larger particle sizes make it difficult to obtain high resolution images.

Many of the disadvantages accompanying the use of dry toner compositions have been avoided by the use of liquid developers or toners. For example, liquid toners contain smaller particles than dry toners resulting in higher resolution images. In addition, liquid toners are not triboelectrically charged; therefore, they are much less sensitive to changes in ambient temperature and humidity. Since the toner particles in a liquid developer are contained within a fluid phase, toner scatter and dust accumulation do not occur within the printer. In addition, the particles being contained within a liquid matrix will not become airborne thus eliminating the risk of inhalation of the particles.

Liquid toners typically comprise an electrically insulating liquid that serves as a carrier for a dispersion of charged particles known as toner particles composed of a colorant and a polymeric binder. A charge control agent is often included as a component of the liquid developer in order to regulate the polarity and magnitude of

the charge on the toner particles. Liquid toners can be categorized into two primary classes, for convenience, the two classes will be referred to as conventional liquid toners and organosol toners.

Of particular utility are the class of liquid toners which make use of self-stable  
5 graft copolymers dispersed in an organic solvent (organosols) as polymeric binders to promote self-fixing of a developed latent image. U.S. Patent Nos. 3,753,760; 3,900,412; 3,991,226; 4,476,210; 4,789,616; 4,728,983; 4,925,766; 4,946,753; 4,978,598 and 4,988,602 describe compositions and uses of graft copolymer organosols. Exemplary liquid electrophotographic, pigmented inks made using self-  
10 stable graft copolymer organosols are illustrated by Kosel in U.S. 3,900,412.

Self-stable organosols are colloidal (0.1-1 micron diameter) particles of polymeric binder that are typically synthesized by nonaqueous dispersion polymerization in a low dielectric hydrocarbon solvent. These organosol particles are sterically-stabilized with respect to aggregation by the use of a physically-adsorbed or  
15 chemically-grafted soluble polymer. Details of the mechanism of such steric stabilization are provided in Napper, D.H., Polymeric Stabilization of Colloidal Dispersions, Academic Press, New York, NY, 1983. Procedures for effecting the synthesis of self-stable organosols are known to those skilled in the art and are described in Dispersion Polymerization in Organic Media, K.E. J. Barrett, ed., John  
20 Wiley: New York, NY, 1975. Although it is generally recognized that the solvation of the particle is critical in the formation of a dispersion, none of the foregoing references recognize the utility of a gel in forming a stable dispersion.

The most commonly used non-aqueous dispersion polymerization method is a free radical polymerization carried out when one or more ethylenically-unsaturated  
25 (typically acrylic) monomers, soluble in a hydrocarbon medium, are polymerized in the presence of a preformed amphipathic polymer. The preformed amphipathic polymer, commonly referred to as the stabilizer, has two distinct ends, one essentially insoluble in the hydrocarbon medium, the other freely soluble. When the polymerization proceeds to a fractional conversion of monomer corresponding to a  
30 critical molecular weight, the solubility limit is exceeded and the polymer precipitates from solution, forming a core particle. The amphipathic polymer then either adsorbs onto or covalently bonds to the core, which core continues to grow as a discrete

particle. The particles continue to grow until monomer is depleted; the adsorbed amphipathic polymer "shell" acts to sterically-stabilize the growing core particles with respect to aggregation. The resulting core/shell polymer particles comprise a self-stable, nonaqueous colloidal dispersion (organosol) comprised of distinct spherical  
5 particles in the size (diameter) range 0.1-0.5 microns.

The resulting organosols can be subsequently converted to liquid toners by simple incorporation of the colorant (pigment) and charge director, followed by high shear homogenization, ball-milling, attritor milling, high energy bead (sand) milling or other means known in the art for effecting particle size reduction in a dispersion.  
10 The input of mechanical energy to the dispersion during milling acts to break down aggregated pigment particles into primary particles (0.05-1.0 micron diameter) and to "shred" the organosol into fragments that adhere to the newly-created pigment surface, thereby acting to sterically-stabilize the pigment particles with respect to aggregation. The charge director may physically or chemically adsorb onto the  
15 pigment, the organosol or both. The result is a sterically-stabilized, charged, nonaqueous pigment dispersion in the size range 0.1-2.0 microns, with typical toner particle diameters between 0.1-0.5 microns. Such a sterically-stabilized dispersion is ideally suited for use in high resolution printing.

Rapid self-fixing is a critical requirement for liquid toner performance to  
20 avoid printing defects (such a smearing or trailing-edge tailing) and incomplete transfer in high speed printing. A description of these types of defects and methods of preventing them using film forming compositions are described in U.S. Patent Nos. 5,302,482; 5,061,583; 4,925,766; 4,507,377; and 4,480,022.

Another important consideration in formulating a liquid toner is the tack of the  
25 image on the final receptor. If the image has a residual tack, then the image may become embossed or picked off when placed in contact with another surface. This is especially a problem when printed sheets are placed in a stack. If the image is tacky, it may transfer to the backside of the adjacent sheet. To address this concern, a film laminate or protective layer is typically placed over the surface of the image. This  
30 adds both extra cost of materials and extra process steps to apply the protective layer.

It is further known in the art that film-forming liquid toners fabricated using self-stable organosols generally exhibit excellent aggregation stability; however, the

sedimentation stability of such inks is poor. Once the components of an organosol ink have settled, they are generally difficult if not impossible to redisperse to a degree of dispersion equivalent to the original, unsettled ink. This situation arises because self-stable organosol inks settle into closely packed, dilatant sediments, and irreversible  
5 film formation can occur in these sediments when the volume fraction of organosol in the sediment exceeds the critical volume fraction required for film formation to occur (generally greater than 70 volume percent organosol). Hence, there is a need for liquid ink compositions that will overcome this poor sedimentation and redispersion behavior of organosol inks.

10 Attempts have been made to provide liquid developers having improved storage and thermal stability using a two component gel/latex system. U.S. Patent Nos. 4,374,918; 4,363,863; 4,306,009; GB 2,066,493; and GB 2,065,320 describe liquid developers incorporating polymers having borderline solubility in the carrier solvent. A weakly crosslinked stabilizing gel and a separate latex (or gelatex) are  
15 used as a dispersant and/or fixative. The gel and latex are separate molecular species in the ink formulation. That is, the weakly crosslinked gel and latex are not covalently bonded to each other. Separate molecular components can lead to preferential depletion of one of the two components during extended printing, thus adversely effecting print quality. Separate components can also result in a broad  
20 molecular weight distribution for stabilizing the gel, which may have an adverse affect on the toner charge characteristics. In addition, the separate materials may lead to high free phase conductivity.

Recent attempts have also been made to overcome the poor sedimentation stability of pigmented liquid toners by replacing the pigment with a dye of  
25 significantly lower density. U.S. 4,816,370 describes liquid developers using a thermally reversible, flocculated, dyed organosol. U.S. 4,476,210 and 4,762,764 similarly describe liquid developers using self-stable dyed organosols. The colorants used in each of these references are dyes rather than pigments. It is well established in the art that dyes are less stable to light and have a tendency to migrate (bleed) or  
30 sublime. Even though dyes have inherent advantages, such as transparency of the colors and less interference with the characteristics of the thermoplastic binders, their poor light stability often times overrides these advantages. Moreover, because of the

typically low mass of dye molecules relative to pigments, the charge/mass of dyed toners is generally higher by an order of magnitude than pigmented toners. This can lead to problems in printing dye-based inks to adequate optical density using liquid electrographic development processes.

5           More recent attempts to address the poor sedimentation and redispersion behavior of pigmented organosol inks are described in U.S. 5,698,616 and U.S. 5,652,282. These patents describe liquid inks that make use of a gel organosol to improve the sedimentation and redispersion properties of pigmented inks. The gel organosols disclosed in these patents are graft copolymers dispersed in an aliphatic hydrocarbon carrier solvent. The gel organosols, comprising graft copolymers in  
10           which a high molecular weight polymeric steric stabilizer (graft stabilizer) is covalently bonded to an insoluble, high molecular weight thermoplastic polymeric core, have the ability to form a three dimensional network of controlled rigidity. The gels are formed by manipulating the solubility parameter difference between the graft  
15           stabilizer and the carrier solvent to a range of 2.5-3.0 MPa<sup>1/2</sup>. The gel structures were found to improve liquid ink performance, particularly pigmented liquid electrographic/electrophotographic ink performance, by increasing the sedimentation stability and redispersability of the colorants without compromising print quality, charge stability, or ink transfer performance.

20           Another issue relating to images printed using liquid inks is image durability. Durability refers to the resistance of the printed images to damage by blocking when printed sheets are stacked, erasure resistance, scratch resistance, and abrasion resistance. Generally, images printed with liquid toners are less durable than images printed using dry toners. The lower durability of liquid ink printed images may result  
25           from a variety of reasons known to those skilled in the art. These include the relatively lower thickness of printed ink films made using liquid toners, the relatively lower adhesive strength of some liquid toners to their print receptors, and the relatively lower cohesive strength of some printed liquid toners due to the comparatively lower glass transition temperature of the polymeric binders used in  
30           these liquid inks.

          Various methods have been proposed to improve the durability of printed images made using liquid toners. In particular, U.S. 5,886,067 describes improved

durability liquid inks comprising a controlled crystallinity organosol that is not a gel organosol. The controlled crystallinity organosol comprises an insoluble core and a soluble graft stabilizer prepared from a side-chain or main chain crystallizing polymeric moiety that independently and reversibly crystallizes at or above 22°C.

- 5 Such controlled crystallinity organosols impart improved blocking resistance, erasure resistance, and abrasion resistance to liquid inks containing the organosol. Suitable crystallizing polymeric moieties for incorporation into a graft stabilizer include  $>C_{14}$  acrylic and methacrylic esters which do not form gel organosols based upon solubility parameter difference between graft stabilizer and carrier solvent.

10

### SUMMARY OF THE INVENTION

- In one aspect, the invention features a rapidly self-fixing ink with improved sedimentation and redispersion characteristics useful as a liquid toner in ionographic or electrographic (electrophotographic or electrostatic) imaging and printing processes. The ink is comprised of a polymeric binder in the form of a graft
- 15 copolymer gel, formed by acid/base interactions, dispersed in an organic solvent or solvent blend having a Kauri-Butanol (KB) number less than 30, and optionally contains one or more colorants. The colorants may take the form of dyes or pigments. One or more charge controlling additives may optionally be added to the ink formulation.

20

- "Kauri-Butanol" refers to an ASTM Test Method D1133-54T. The Kauri-Butanol Number (KB) is a measure of the tolerance of a standard solution of kauri resin in 1-butanol to an added hydrocarbon diluent and is measured as the volume in milliliters (ml) at 25°C of the solvent required to produce a certain defined degree of turbidity when added to 20 g of a standard kauri-1-butanol solution. Standard values
- 25 are toluene (KB = 105) and 75% by volume of heptane with 25% by volume toluene (KB = 40). There is an approximately linear relationship between the Hildebrand solubility parameter and the KB number for hydrocarbons: Hildebrand Solubility Parameter ( $MPa^{1/2}$ ) =  $2.0455[6.3 + 0.03KB \text{ (ml)}]$ .

- The graft copolymer is comprised of a soluble or marginally insoluble high
- 30 molecular weight (co)polymeric steric stabilizer ("graft stabilizer") covalently bonded to an insoluble, high molecular weight thermoplastic (co)polymeric backbone



("organosol core"). The gel is formed by incorporating a low percentage of each of two monomer types, acidic and basic monomers, in the graft stabilizer, in the organosol core, or in both the graft stabilizer and organosol core. The acidic and basic monomers form weak, reversible intermolecular non-covalent bonds resulting from acid/base interactions between dispersed graft copolymer particles, thereby forming a gel organosol.

The gel organosols provide a new approach to improving the sedimentation and redispersion properties of pigmented inks. The method of inducing gelation does not require manipulation of the relative difference in solubility parameter between the graft stabilizer and the carrier solvent into a range (solubility parameter difference greater than  $2.5 \text{ MPa}^{1/2}$ ) that acts to reduce agglomeration stability of the graft copolymer. This allows the ink formulator increased flexibility in selection of monomer components of the graft stabilizer, as well as greater flexibility in carrier fluid selection.

For example, side-chain crystallizable monomers that have a high degree of solubility in the carrier solvent may be incorporated into the graft stabilizer without sacrificing gelation characteristics. The use of crystallizable polymeric moieties to improve the durability of non-gel organosol inks has been disclosed in U.S. 5,886,067. Heretofore, the use of such crystallizable polymeric moieties at high weight percentages in a graft stabilizer has prevented the formation of gel organosols owing to the relative solubility parameter difference between the graft stabilizer and the carrier solvent falling in the range of good solubility ( $0-2.5 \text{ MPa}^{1/2}$ ). It would be advantageous to combine the characteristics of a gel organosol and a controlled-crystallinity organosol into a single graft copolymer composition.

The acid-base gel organosols are particularly useful for their ability to form a three dimensional gel of controlled rigidity that can be reversibly reduced to a fluid state by shearing or heating the organosol or an ink composition containing the organosol. The gels impart useful properties to the liquid ink, notably improved sedimentation stability of the colorant, without compromising print quality or ink transfer performance. The inks formulated with the gels also exhibit improved redispersion characteristics upon settling, and do not form dilatant sediments such as those formed by non-gelled organosol inks. These characteristics of gel inks facilitate

preparation and use of high solids ink concentrates (greater than 2% by weight solids, more preferably greater than 10% by weight solids), thus providing an increased number of printed pages or images from a given volume of ink.

The liquid inks will be described below with respect to electrophotographic printing; however, it is to be understood that these liquid inks are not limited in their utility and may also be employed in other electrographic printing processes, high speed printing presses, photocopying apparatus, microfilm reproduction devices, facsimile printing, ink jet printer, instrument recording devices, and the like.

### DETAILED DESCRIPTION

A liquid ink composition is provided comprising a colorant and an gel organosol dispersed in a liquid or liquid blend having a Kauri-Butanol (KB) number less than 30. The liquid ink composition is resistant to sedimentation and is capable of rapidly self-fixing, which is particularly useful in electrophotographic, ionographic, or electrostatic imaging and conventional printing processes.

The gel is an amphipathic copolymeric gel comprised of a soluble or marginally insoluble high molecular weight (co)polymeric steric stabilizer which is non-covalently crosslinked by acid-base interactions and covalently bonded to an insoluble, thermoplastic (co)polymeric core. The covalently bonded graft steric stabilizer is non-covalently crosslinked by acid-base interactions to such an extent that it behaves as an extremely high molecular weight copolymer near its incipient phase separation point in the dispersant liquid. The non-covalently crosslinked graft stabilizer remains in a freely flowing, easily handled solution until the graft stabilizer is covalently bonded to the insoluble core, at which point a gel organosol is formed.

Gel organosols are dispersions in which the attractive interactions between the elements of the dispersed phase are so strong that the whole system develops a rigid network structure and, under small stresses, behaves elastically. The characteristic of organosol gelation is visibly apparent to one skilled in the art. The non-covalently crosslinked acid-base gel organosols rapidly gel to form a voluminous polymer sediment and a substantially clear supernatant layer of carrier liquid upon standing.

Gelation of the graft copolymer organosol is induced by forming non-covalent bonds between a plurality of graft stabilizer chains using weak, reversible acid-base

interactions between an acidic polymerizable compound and a basic polymerizable compound. The acidic or basic polymerizable compounds may both be polymerized into the organosol shell, or one of the polymerizable compounds may be polymerized into the shell, and the other in the core. It is also possible to prepare two separate and distinct graft copolymer compositions, one containing an acidic or basic polymerizable compound in the core or shell, the other containing the basic or acidic polymerizable compound in the shell. In this embodiment, gelation would not occur until the two organosols were blended together.

The strength of the gel (and hence sedimentation resistance of the ink) can be readily manipulated by controlling the extent to which the graft stabilizer is non-covalently crosslinked. Greater gel strength (greater sedimentation resistance) is obtained by increasing the crosslink density (percentage of acidic or basic polymerizable compound) in the graft copolymer.

Superior stability of the dispersed toner particles with respect to aggregation is obtained when at least one of the polymers or copolymers (denoted as the stabilizer) is an amphipathic substance containing at least one oligomeric or polymeric component of molecular weight at least 500 that is solvated by the carrier liquid. In other words, the selected stabilizer, if present as an independent molecule, would have some finite solubility in the carrier liquid.

The solubility of a material in a given solvent may be predicted from the absolute difference in Hildebrand solubility parameter of the solute relative to the solvent. The solutes will exist as true solutions or in a highly solvated state when the absolute difference in Hildebrand solubility parameter is less than approximately  $1.5 \text{ MPa}^{1/2}$ . When the absolute difference in Hildebrand solubility parameter exceeds approximately  $3.0 \text{ MPa}^{1/2}$ , the solute will phase separate from the dispersant, forming a solid, insoluble, non-flowing, non-gelled mass. Those solutes having an absolute difference in Hildebrand solubility parameters between  $1.5 \text{ MPa}^{1/2}$  and  $3.0 \text{ MPa}^{1/2}$  are considered to be weakly solvated or marginally insoluble.

The solubility parameters of the graft stabilizers, as well as the carrier liquids in which the stabilizers are dispersed, are calculated using values for the Hildebrand solubility parameter of the monomers used to prepare the graft stabilizer and the carrier liquids obtained using the group contribution method developed by Small, P.

A., J. Appl. Chem., **3**, 71 (1953) using Small's group contribution values listed in Table 2.2 on page VII/525 in the Polymer Handbook, 3rd Ed., J. Brandrup & E. H. Immergut, Eds. John Wiley, NY, pp 519-557 (1989). The Hildebrand solubility parameter for a copolymer may be calculated using a volume fraction weighting of the individual Hildebrand solubility parameters for each monomer comprising the copolymer. Similarly, the Hildebrand solubility parameter for a mixture may be calculated using a volume fraction weighting of the individual Hildebrand solubility parameters for each component of the mixture. Thus, the Hildebrand solubility parameter for a mixture of solvents or polymerizable compounds may be calculated using a volume fraction weighting of the individual Hildebrand solubility parameters for each chemical compound comprising the solvent mixture.

Table 1 lists the Hildebrand solubility parameters for some common carrier liquids used in an electrophotographic toner and the Hildebrand solubility parameters and glass transition temperatures for some common monomers used in synthesizing organosols.

**Table I Hildebrand Solubility Parameters**

**Solvent Values at 25°C**

| Solvent<br>Commercial Trade<br>Designation | Kauri-Butanol Number<br>by ASTM Method<br>D1133-54T(ml) | Hildebrand Solubility<br>Parameter (MPa <sup>1/2</sup> ) |
|--|---|--|
| NORPAR 15                                  | 18  | 13.99  |
| NORPAR 13                                  | 22  | 14.24  |
| NORPAR 12                                  | 23  | 14.30  |
| ISOPAR V                                   | 25  | 14.42  |
| EXXSOL D80                                 | 28  | 14.60  |

Source: Calculated from equation #31 of Polymer Handbook, 3rd Ed., J. Brandrup E.H. Immergut, Eds. John Wiley, NY, p. VII/522 (1989).

**Monomer Values at 25°C**

| <b>Monomer Name</b>       | <b>Hildebrand Solubility<br/>Parameter (MPa<sup>1/2</sup>)</b> | <b>Glass Transition<br/>Temperature (°C)*</b> |
|---------------------------|--|---|
| n-Octadecyl Methacrylate  | 16.77  | -100  |
| n-Octadecyl Acrylate      | 16.82  | -55   |
| Lauryl Methacrylate       | 16.84  | -65   |
| Lauryl Acrylate           | 16.95  | -30   |
| 2-Ethylhexyl Methacrylate | 16.97  | -10   |
| 2-Ethylhexyl Acrylate     | 17.03  | -55   |
| n-Hexyl Methacrylate      | 17.13  | -5  |
| t-Butyl Methacrylate      | 17.16  | 107   |
| n-Butyl Methacrylate      | 17.22  | 20  |
| n-Hexyl Acrylate          | 17.30  | -60   |
| n-Butyl Acrylate          | 17.45  | -55   |
| Ethyl Methacrylate        | 17.90  | 66  |
| Ethyl Acrylate            | 18.04  | -24   |
| Methyl Methacrylate       | 18.17  | 105   |
| Vinyl Acetate             | 19.40  | 30  |
| Methyl Acrylate           | 20.2   | 5   |

Calculated using Small's Group Contribution Method, Small, P.A. Journal of Applied Chemistry 3 p. 71 (1953). Using Group Contributions from Polymer Handbook, 3rd Ed., J. Brandrup E.H. Immergut, Eds., John Wiley, NY, p. VII/ 525 (1989).

\* Polymer Handbook, 3rd Ed., J. Brandrup E.H. Immergut, Eds., John Wiley, NY, pp. VII/209-277 (1989).

The carrier liquid may be selected from a wide variety of materials that are known in the art, but preferably has a Kauri-Butanol number less than 30. The liquid is typically oleophilic, chemically stable under a variety of conditions, and electrically insulating. Electrically insulating refers to a liquid having a low dielectric

constant and a high electrical resistivity. Preferably, the liquid has a dielectric constant of less than 5, more preferably less than 3. Electrical resistivities of carrier liquids are typically greater than  $10^9$  Ohm-cm, more preferably greater than  $10^{10}$  Ohm-cm. The carrier liquid must also be relatively nonviscous to allow movement of the charged particles during development, and sufficiently volatile to permit its removal from the final imaged substrate, but sufficiently non-volatile to minimize evaporative losses in the developer. In addition, the carrier liquid should be chemically inert with respect to the materials or equipment used in the liquid electrophotographic process, particularly the photoreceptor and its release surface.

- Examples of suitable carrier liquids include aliphatic hydrocarbons (n-pentane, hexane, heptane and the like), cycloaliphatic hydrocarbons (cyclopentane, cyclohexane, and the like), aromatic hydrocarbons (benzene, toluene, xylene, and the like), halogenated hydrocarbon solvents (chlorinated alkanes, fluorinated alkanes, chlorofluorocarbons, and the like), silicone oils, and blends of these solvents.
- Preferred carrier liquids include branched paraffinic solvent blends such as those available under the trade designations ISOPAR G, ISOPAR H, ISOPAR K, ISOPAR L, ISOPAR M and ISOPAR V (all available from Exxon Corporation, NJ), and most preferred carriers are the aliphatic hydrocarbon solvent blends such as those available under the trade designations NORPAR 12, NORPAR 13 and NORPAR 15 (all available from Exxon Corporation, NJ).

The organosol is a graft copolymer prepared by chemically bonding a generally soluble (co)polymer to a generally insoluble (co)polymer resin core. Any number of reactions known to those skilled in the art may be used to effect grafting of the soluble polymeric stabilizer to the organosol core during free radical polymerization. Common grafting methods include random grafting of polyfunctional free radicals; ring-opening polymerizations of cyclic ethers, esters, amides or acetals; epoxidations; reactions of hydroxyl or amino chain transfer agents with terminally-unsaturated end groups; esterification reactions (e.g., glycidyl methacrylate undergoes tertiary-amine catalyzed esterification with methacrylic acid); and condensation polymerization.

The composition of the graft stabilizer is normally selected such that the Hildebrand Solubility Parameter of the graft stabilizer (shell) closely matches that of

the carrier liquid in order to ensure that the stabilizer will be sufficiently solvated to dissolve in the carrier solvent. Virtually any polymerizable compound that exhibits a Hildebrand Solubility Parameter difference less than  $3.0 \text{ MPa}^{1/2}$  relative to the carrier liquid may be used in forming a graft stabilizer. In addition, polymerizable

5 compounds that exhibit a Hildebrand Solubility Parameter difference greater than  $3.0 \text{ MPa}^{1/2}$  relative to the carrier liquid may be used in forming a copolymeric graft stabilizer, provided that the effective Hildebrand Solubility Parameter difference for the stabilizer is less than  $3.0 \text{ MPa}^{1/2}$  relative to the carrier liquid. The absolute difference in Hildebrand Solubility Parameter between the graft stabilizer (shell) and

10 the carrier liquid is preferably less than  $2.6 \text{ MPa}^{1/2}$ .

Preferred polymerizable compounds useful in forming the graft stabilizer are the  $\text{C}_6\text{-C}_{30}$  acrylic and methacrylic esters. Examples of suitable polymerizable compounds for use in the graft stabilizer composition include monomers such as hexyl acrylate, 2-ethylhexyl acrylate, decyl acrylate, dodecyl (lauryl) acrylate,

15 octadecyl (stearyl) acrylate, behenyl acrylate, hexyl methacrylate, 2-ethylhexyl(methacrylate), decyl acrylate, dodecyl (lauryl) methacrylate, octadecyl (stearyl) methacrylate, and other alkyl acrylates and methacrylates.

Preferrably, the polymerizable compounds are also crystallizable compounds having crystallization (melting) temperatures above room temperature ( $25^\circ\text{C}$ ). Such

20 crystallizable, polymerizable compounds are particularly useful in forming graft stabilizers that yield organosols and inks exhibiting improved image durability, including blocking and erasure resistance. Crystallizable, polymerizable compounds suitable for incorporation into a graft stabilizer include  $>\text{C}_{14}$  acrylic and methacrylic esters. Preferred crystallizable, polymerizable compounds include octadecyl acrylate

25 and behenyl acrylate.

Other monomers, macromers, or polymers may be used either alone or in conjunction with the aforementioned materials, including melamine and melamine formaldehyde resins, phenol formaldehyde resins, epoxy resins, polyester resins, styrene and styrene/acrylic copolymers, acrylic and methacrylic esters, cellulose acetate and cellulose acetate-butyrate copolymers, and poly(vinyl butyral)

30 copolymers.

Preferred molecular weight ranges for the graft stabilizer are 5,000-1,000,000 Daltons (Da), more preferably = 50,000-500,000 Da, most preferably = 150,000-250,000 Da. The polydispersity of the graft stabilizer also has an affect on imaging and transfer performance of the liquid toners. Generally, it is desirable to maintain the polydispersity (the ratio of the weight-average molecular weight to the number average molecular weight) of the graft stabilizer below 15, more preferably below 5, most preferably below 2.5.

Gelation is induced as a result of to weak attractive forces arising from the electron donor-acceptor interaction between a Lewis acid and a Lewis base, as described in Allan F. M. Barton, Handbook of Solubility and Other Cohesion Parameters (CRC Press: Boca Raton, FL, 1991 pp. 71-72). In one embodiment, gelation is induced by incorporating both a Lewis acid-functional and a Lewis base-functional polymerizable compound in the graft copolymer organosol. The Lewis acid and base-functional polymerizable compounds may be incorporated in the organosol core, in the organosol shell (graft stabilizer), or in both the organosol core and shell. Preferably, the Lewis acid or base-functional polymerizable compound is incorporated into the organosol core, and the corresponding Lewis acid or base-functional polymerizable compound needed to form the Lewis acid/base pair is incorporated into the organosol shell (graft stabilizer).

In another embodiment, two distinct organosols are prepared, one with an acid-functional polymerizable compound in the core and/or shell, and a separate graft copolymer organosol with basic-functional polymerizable compound in the core and/or shell. This approach has the advantage that gelation will not occur until the two organosols are blended together, or until inks comprising the two distinct organosols are blended together. This permits easy handling of two self-stable organosols until the organosols are combined with a colorant to make a gel ink.

Suitable Lewis acid polymerizable compounds include acrylic acid, 2-acrylamido-2-methyl propane sulfonic acid, crotonic acid, itaconic acid, maleic acid, methacrylic acid, pentaerythritol dimethacrylate, and 4-vinyl benzoic acid.

Suitable Lewis base polymerizable compounds include allyl alcohol, allyl amine, allyl diethylamine, allyl ethylamine, allyl dimethylamine, allyl hydroxyethyl ether, N-allyl piperidine, p-amino styrene, diallyl amine, bis-diallylamino methane, t-butylamino methacrylate, diethylaminoethyl methacrylate, diallyl methylamine, N,N-



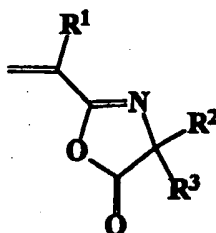
diallylmelamine, 2-dimethylaminoethyl methacrylate, dimethylaminopropyl acrylamide, 2-dimethylamino methyl styrene, 3-dimethylamino neopentyl acrylate, 2,3-dihydroxy propyl acrylate, 2-diisopropylaminoethyl methacrylate, 4-hydroxybutyl acrylate, 4-hydroxybutyl methacrylate, 2-hydroxyethyl acrylate, 2-hydroxyethyl methacrylate, 2-hydroxypropyl acrylate, 2-hydroxypropyl methacrylate, 4-hydroxy styrene, vinyl benzene alcohol, vinyl benzene dimethylamine, 2-vinyl pyridine, 4-vinyl pyridine, and N-vinyl-2-pyrrolidone.

The extent of gelation of the acid-base gel organosol may be controlled by manipulating the concentration of acidic or basic polymerizable compound incorporated into the graft copolymer. Generally, a higher concentration of acidic or basic polymerizable compound leads to a higher apparent crosslink density and therefore a stronger gel. However, too high a concentration of acidic or basic polymerizable monomer will cause the graft copolymer to solidify into a very high molecular weight, solid-like polymer that is not suitable for incorporation into a crosslinked gel organosol. Preferably, 0.1-5.0% w/w, more preferably 0.5-4.0% w/w, most preferably 1.0-3.0% w/w of each polymerizable compound (acid or base) is used. However, the preferred concentration range of acidic or basic polymerizable compound will vary somewhat depending on the specific acidic or basic polymerizable compound. Preferably, the relative amount of Lewis acid polymerizable compound is stoichiometrically matched by the amount of Lewis base polymerizable compound.

As discussed above, the organosol is a graft copolymer dispersion formed by covalently bonding the graft stabilizer to an insoluble resin core. The grafting reaction generally occurs between a grafting site incorporated into the graft stabilizer and a reaction site in the polymerizing or polymerized core. Preferably, the grafting reaction proceeds by reaction of an isocyanate to a hydroxyl group to form a urethane linkage between the stabilizer and the core. The grafting site is preferably formed by incorporating hydroxyl groups into the graft stabilizer during a first free radical polymerization and catalytically reacting all or a portion of these hydroxyl groups with an ethylenically unsaturated aliphatic isocyanate (e.g. meta-isopropenyldimethylbenzyl isocyanate [TMI] or isocyanatoethylmethacrylate [IEM]) to form a polyurethane linkage during a subsequent non-free radical reaction step. The graft stabilizer is then covalently bonded to the nascent insoluble acrylic

(co)polymer core via reaction of the unsaturated vinyl group of the grafting site with ethylenically-unsaturated core monomers (e.g. vinyl esters, particularly acrylic and methacrylic esters with carbon numbers < 6 or vinyl acetate; vinyl aromatics, such as styrene; acrylonitrile; n-vinyl pyrrolidone; vinyl chloride and vinylidene chloride) during a subsequent free radical polymerization step.

Other methods of effecting grafting of the preformed polymeric stabilizer to the incipient insoluble core particle are known to those skilled in the art. For example, alternative grafting protocols are described in sections 3.7-3.8 of Barrett Dispersion Polymerization in Organic Media, K. E. J. Barrett, ed., (John Wiley: New York, 1975), pp. 79-106. A particularly useful method for grafting the polymeric stabilizer to the core utilizes an anchoring group. The function of the anchoring groups is to provide a covalent link between the core part of the particle and the soluble component of the steric stabilizer. Suitable monomers containing anchoring groups include: adducts of alkenylazlactone comonomers with an unsaturated nucleophile containing hydroxy, amino, or mercaptan groups, such as 2-hydroxyethylmethacrylate, 3-hydroxypropylmethacrylate, 2-hydroxyethylacrylate, pentaerythritol triacrylate, 4-hydroxybutylvinylether, 9-octadecen-1-ol, cinnamyl alcohol, allyl mercaptan, methallylamine; and azlactones, such as 2-alkenyl-4,4-dialkylazlactone of the structure



where  $R^1 = H$ , or alkyl having 1 to 5 carbons, preferably one carbon,  $R^2$  and  $R^3$  are independently lower alkyl groups having 1 to 8 carbons, preferably 1 to 4 carbons.

Most preferably, however, the grafting mechanism is accomplished by grafting an ethylenically-unsaturated isocyanate (e.g., dimethyl-m-isopropenyl benzyliisocyanate, available from American Cyanamid) to hydroxyl groups previously incorporated into the graft stabilizer precursor (e.g., hydroxy ethyl methacrylate).

The insoluble organosol core is the dispersed phase of the graft copolymer dispersion. The core polymer is generally made *in situ* by copolymerization with the stabilizer monomer. The solubility parameter of the core is generally chosen such that it differs substantially from that of the dispersion medium in order to ensure that the core monomers will phase separate during dispersion polymerization (forming the core). Virtually any monomer or combination of monomers may be used in forming the organosol core, provided that the core as a whole is insoluble in the carrier liquid. Preferably, the Hildebrand solubility parameter difference between the core as a whole and the carrier liquid exceeds  $3.0 \text{ MPa}^{1/2}$ . Examples of polymerizable compounds suitable for use in the organosol core include  $\text{C}_1\text{-C}_5$  acrylate and methacrylate esters such as, methyl acrylate, ethyl acrylate, butyl acrylate, methyl methacrylate, ethyl methacrylate), butyl methacrylate, styrene, and vinyl acetate.

Monomers that individually are very soluble in the dispersant medium may also be incorporated into the core in small amounts with monomers which, when polymerized, are not very soluble in the carrier liquid. The effect of incorporating a monomer having good solubility in the carrier liquid into the organosol core will generally be to cause carrier liquid absorption or swelling of the core. This can be particularly useful when higher core  $T_g$ 's are desired, for example, to improve image durability, but when rapid self-fixing of the ink is still required in the imaging process. By increasing the affinity of the core for the carrier liquid, the carrier liquid will be imbibed into the core and may act to plasticize the core, permitting rapid self-fixing of an ink comprising a high core  $T_g$  organosol even at temperatures below the normal minimum film-forming temperature.

Preferably, the Hildebrand solubility parameter difference between the core as a whole and the carrier liquid exceeds  $3.0 \text{ MPa}^{1/2}$ . The amount of soluble monomer incorporated into the core typically ranges between 1-30% w/w, more preferably 5-20% w/w. Examples of soluble polymerizable compounds suitable for use in the organosol core include  $\text{C}_6\text{-C}_{30}$  acrylate and methacrylate esters such as cyclohexyl acrylate, hexyl acrylate, 2-ethylhexyl acrylate, decyl acrylate, dodecyl (lauryl) acrylate, octadecyl (stearyl) acrylate, behenyl acrylate, cyclohexyl methacrylate, hexyl methacrylate, 2-ethylhexyl(methacrylate), decyl acrylate, dodecyl (lauryl)

methacrylate, octadecyl (stearyl) methacrylate, and other acrylates and methacrylates which meet the solubility parameter requirements described above.

In selecting polymerizable compounds for use in the organosol core, it is necessary to consider not only their effect on solubility parameter difference between the core and the carrier liquid, but also their effect on the effective glass transition temperature of the core and the resulting impact on imaging performance. Preferably, the  $T_g$  of the core is  $<30^\circ\text{C}$  to allow an ink composition containing the resin as a major component to undergo rapid film formation (rapid self-fixing) in printing or imaging processes carried out at temperatures greater than the core  $T_g$ , preferably at or above  $23^\circ\text{C}$ . Rapid self-fixing assists in avoiding printing defects (such as smearing or trailing-edge tailing) and incomplete transfer in high speed printing. The use of low  $T_g$  thermoplastic polymeric binders to promote film formation is described in Z.W. Wicks, *Film Formation*, Federation of Societies for Coatings Technologies, p 8 (1986).

The  $T_g$  can be calculated for a (co)polymer using known values for the high molecular weight homopolymers (Table I) and the Fox equation expressed below:

$$1/T_g = w_1/T_{g1} + w_2/T_{g2} + w_3/T_{g3} + \dots$$

where  $w_i$  is the weight fraction of monomer "i" and  $T_{gi}$  is the glass transition temperature of the high molecular weight homopolymer of monomer "i" as described in Wicks, A.W., F.N. Jones & S.P. Pappas, *Organic Coatings*, 1, John Wiley, NY, pp 54-55 (1992).

Another reason for using polymer particles in which the core has a  $T_g < 30^\circ\text{C}$  is that these particles can coalesce into a resinous film at room temperature. The overprinting capability of a toner is related to the ability of the polymer particles to deform and coalesce into a resinous film during the air drying cycle of the electrophoretically deposited toner particles. The coalescent particles permit the electrostatic latent image to discharge during the imaging cycle so another image can be overprinted. On the other hand, non-coalescent particles retain their shape even after being air dried on the photoreceptor. The points of contact are then few compared to a homogeneous or continuous film-forming latex, and as a result, some of the charges are retained on the unfused particles, repelling the next toner.

A toner layer made of polymer particles having a core with a  $T_g > 30^\circ\text{C}$  may be made to coalesce into a film at room temperature if the stabilizer/core ratio is high enough. Thus the choice of stabilizer/(core + stabilizer) ratios in the range 20 wt.% to 80 wt.% can give coalescence at room temperature with core  $T_g$  values in a corresponding range  $25^\circ\text{C}$  to  $105^\circ\text{C}$ . With a core  $T_g < 30^\circ\text{C}$  the preferred range of stabilizer/(core + stabilizer) ratio is 10 to 40 wt.%.

The integrity of the toned image during partial removal of the solvent also depends upon the core  $T_g$ , with lower  $T_g$  promoting film strength and image integrity at the cost of additional image tack. An organosol core  $T_g$  below room temperature is preferred to ensure that the toner will in fact film form. Preferably, the minimum film forming temperatures are between about  $5-45^\circ\text{C}$  and the organosol core  $T_g$  is below room temperature to allow the toner to form a film and maintain good image integrity during solvent removal and good cohesive strength during image transfer from the photoconductor onto either a transfer medium or receptor.

To yield the best offset transfer efficiency without causing excessive residual tack and blocking of the transferred image, an organosol core having a  $T_g$  between  $-10$  and  $20^\circ\text{C}$  is preferred, more preferably between  $-5$  and  $15^\circ\text{C}$ . Under constant transfer roll pressure conditions, lower core  $T_g$  organosols exhibit 100% transfer at lower temperatures than high  $T_g$  organosols. The results of toner transfer testing are dependent upon the extent of self-fixing or dryness of the toner film. In addition, toned images will require some finite drying time in order to allow film formation to occur. This drying can be accelerated by using heated air, vacuum drying, an electrostatically-biased or unbiased squeegee (to hydraulically remove excess dispersant) or other similar methods known in the art, e.g., the system described in U.S. Patent No. 5,300,990. The rate of film formation (self-fixing) may also be accelerated by adding a plasticizer to the toned image to effectively lower the minimum film forming temperature.

Residual image tack after transfer may be adversely affected by the presence of high tack monomers, such as ethyl acrylate, in the organosol. In addition, the cohesive strength and durability of the ink film generally decreases as the organosol core  $T_g$  decreases. Ideally, organosol core  $T_g$  would be selected to be as high as

possible in order to reduce residual tack and maximize image durability. However, this is often not practical, particularly for liquid inks used in imaging processes in which the inks are required to form a film at room temperature. Consequently, organosols are generally formulated such that the organosol core preferably has a  
5 glass transition temperature ( $T_g$ ) less than high room temperature ( $30^\circ\text{C}$ ) but greater than  $-10^\circ\text{C}$ .

The organosol core comprises approximately 30-95% of the organosol on a weight basis. Thus, the core's  $T_g$  will typically dominate over the stabilizer's  $T_g$  and the organosol  $T_g$  may be taken as a first approximation to be the core  $T_g$ . As  
10 illustrated supra, a simple algebraic method based upon a composition-weighted sum of inverse  $T_g$ 's (using absolute temperatures) can be used to calculate the effective  $T_g$  of a copolymer blend used to make up an organosol core. Similar arguments allow an assumption that the solubility parameter of the stabilizer will generally control the stability of the organosol. Therefore, the stabilizer solubility parameter is preferably  
15 selected to closely match that of the dispersant for maximum aggregation stability.

Using the data in Table 1 and applying the above criteria, it is noted that polymerizable compounds such as methyl acrylate, ethyl acrylate, and vinyl acetate are most suitable for incorporation into an organosol core comprising a single polymerizable compound exclusive of the grafting compound ( $T_g$  between  $-30$  and  
20  $30^\circ\text{C}$  and solubility parameter difference relative to the solvents NORPAR 12 greater than  $3.0 \text{ MPa}^{1/2}$ ). Preferably, the core comprises a copolymer of at least one high glass transition temperature polymerizable compound, such as methyl methacrylate, ethyl methacrylate or butyl methacrylate, and at least one low glass transition polymerizable compound such as ethyl acrylate or butyl acrylate. This allows  
25 variation of the core  $T_g$  over a wide range between  $T_g$ 's of the respective polymerizable compounds by simple variation of the relative weight ratios of the high and low  $T_g$  polymerizable compounds, as calculated using the Fox Equation.

The preferred polymerizable compounds for use in making a copolymer core in NORPAR 12 carrier liquid are methyl methacrylate and ethyl acrylate since these  
30 monomers individually or in any weight ratio exceed the  $3.0 \text{ MPa}^{1/2}$  Hildebrand solubility parameter difference relative to NORPAR 12 and will therefore readily

form an insoluble core in this carrier liquid upon polymerization. In addition, ethyl acrylate and methyl methacrylate permit the variation of the core  $T_g$  anywhere within the range  $-24^{\circ}\text{C}$  to  $105^{\circ}\text{C}$  merely by varying the relative weight ratios of the two monomers in the core.

5           A preferred organosol core composition contains about 75 weight percent ethyl acrylate and 25 weight percent methyl methacrylate, yielding a calculated core  $T_g$  of  $-1^{\circ}\text{C}$ . This permits the toners to rapidly self-fix under normal room temperature or higher development conditions and also produce tack-free fused images that resist blocking at room temperature.

10           Other polymers which may be used either alone or in conjunction with the aforementioned materials include melamine and melamine formaldehyde resins, phenol formaldehyde resins, epoxy resins, polyester resins, styrene and styrene/acrylic copolymers, acrylic and methacrylic esters, cellulose acetate and cellulose acetate-butyrate copolymers, and poly(vinyl butyral) copolymers.

15           If the core/shell ratio is too high, there may be insufficient graft stabilizer present to sterically stabilize the organosol with respect to aggregation. If the core/shell ratio is too low, the polymerization may have insufficient driving force to form a distinct particulate phase, resulting in a copolymer solution, not a self-stable organosol dispersion. The optimal weight ratio of the resin core to the stabilizer shell  
20 is on the order of 1/1 to 15/1, preferably between 2/1 and 10/1, and most preferably between 4/1 and 8/1.

          The particle size of the organosol also influences the imaging, drying, and transfer characteristics of the liquid inks. Preferably, the primary particle size (determined with dynamic light scattering) of the organosol is between about 0.05 and  
25 5.0 microns, more preferably between 0.15 and 1 micron, most preferably between 0.20 and 0.50 microns.

          A liquid ink utilizing the aforementioned gel organosol comprises colorant particles embedded in the thermoplastic organosol resin. Useful colorants are well known in the art and include materials such as dyes, stains, and pigments. Preferred  
30 colorants are pigments that may be incorporated into the polymer resin, are nominally insoluble in and nonreactive with the carrier liquid, and are useful and effective in

making visible the latent electrostatic image. Examples of suitable colorants include: phthalocyanine blue (C.I. Pigment Blue 15:1, 15:2, 15:3, 15:4 and 16), monoarylide yellow (C.I. Pigment Yellow 1, 3, 65, 73 and 74), diarylide yellow (C.I. Pigment Yellow 12, 13, 14, 17 and 83), arylamide (Hansa) yellow (C.I. Pigment Yellow 10, 97, 105 and 111), Pigment Yellow 138, azo red (C.I. Pigment Red 3, 17, 22, 23, 38, 48:1, 48:2, 52:1, 81, 81:1, 81:2, 81:3 and 179), quinacridone magenta (C.I. Pigment Red 122, 202 and 209) and black pigments such as finely divided carbon (Cabot Monarch 120, Cabot Regal 300R, Cabot Regal 350R, Vulcan X72) and the like.

For some applications, it is desirable to use the organosol without an added colorant (dye or pigment) to provide a clear protective overcoat for an underlying image on a permanent receptor. In such cases, the transparent ink may be applied, for example, using well-known coating or electrographic development processes, onto either a temporary imaging receptor or permanent imaging receptor. In the event that the transparent organosol is applied to a permanent image receptor, the organosol should be coated onto the surface of the receptor and any underlying image in order to perform as a protective overcoat. In the event that the organosol is applied to a temporary image receptor, consideration must be given to the reversal of layers that occurs during offset transfer processes. Thus, it may be necessary to coat or develop the transparent organosol as the first layer on a photoreceptive element upon which a multi-colored image is constructed in order to insure that the transparent organosol acts as a protective topcoat upon offset transfer of the image to a permanent image receptor.

The optimal weight ratio of resin (organosol) to colorant in the toner particles is on the order of 1/1 to 20/1, preferably between 3/1 and 10/1, and most preferably between 5/1 and 8/1. The total dispersed material in the carrier liquid typically represents 0.5 to 70 weight percent, preferably between 1 and 25 weight percent, most preferably between 2 and 17 weight percent of the total liquid developer composition.

The gel organosols have been used to fabricate liquid electrophotographic toners that exhibit excellent imaging characteristics in liquid immersion development. For example, the gel organosol liquid toners exhibit low bulk conductivity, low free phase conductivity, and low charge/mass and high mobility, all of which are desirable characteristics for producing high resolution, background-free images with high



optical density. In particular, the low bulk conductivity, low free phase conductivity, and low charge/mass of the toners allow them to achieve high developed optical density over a wide range of solids concentrations, thus improving their extended printing performance relative to conventional toners. In addition, color liquid toners  
5 made based upon these gel organosols on development form transparent films that transmit incident light, consequently allowing the photoconductor layer to discharge.

The toners have low  $T_g$  values with respect to most available toner materials. This enables the toners to form films at room temperature. It is not necessary for any specific drying procedures or heating elements to be present in the apparatus. Normal  
10 room temperature (19-20°C) is sufficient to enable film forming, as is the ambient internal temperature of the apparatus during operation which tends to be at a higher temperature (e.g., 25-40°C) even without specific heating. It is possible to have the apparatus operate at an internal temperature of 40°C or less at the toning station and immediately thereafter where a fusing operation would ordinarily be located.

15 The gel organosol liquid toners also exhibit improved transfer characteristics relative to conventional inks used in the art, particularly with offset transfer processes. The rapid-fixing characteristics of the toners permit their use in liquid development/dry adhesive offset transfer imaging processes, such as the processes described in U.S. 5,650,253 and 5,916,718. Dry adhesive transfer eliminates the need  
20 for coronas or other charging devices to electrostatically assist transfer of a wet image from the imaging surface to the preferred substrate. In addition, dry adhesive transfer of the toner further reduces carry-out of excessive solvent vapors with the image, as would occur with conventional electrostatically-assisted transfer processes.

An electrophotographic liquid toner may be formulated by incorporating a  
25 charge control agent into the liquid ink. The charge control agent, also known as a charge director, provides uniform charge polarity of the toner particles. The charge director may be incorporated into the toner particles using a variety of methods such as chemically reacting the charge director with the toner particle, chemically or physically adsorbing the charge director onto the toner particle (resin or pigment), or  
30 chelating the charge director to a functional group incorporated into the toner particle. A preferred method is via a functional group built into the graft stabilizer. The charge director imparts an electrical charge of selected polarity onto the toner particles. Any

number of charge directors described in the art may be used. For example, the charge director may be introduced in the form of metal salts consisting of polyvalent metal ions and organic anions as the counterion. Suitable metal ions include Ba(II), Ca(II), Mn(II), Zn(II), Zr(IV), Cu(II), Al(III), Cr(III), Fe(II), Fe(III), Sb(III), Bi(III), Co(II),  
5 La(III), Pb(II), Mg(II), Mo(III), Ni(II), Ag(I), Sr(II), Sn(IV), V(V), Y(III), and Ti(IV). Suitable organic anions include carboxylates or sulfonates derived from aliphatic or aromatic carboxylic or sulfonic acids, preferably aliphatic fatty acids such as stearic acid, behenic acid, neodecanoic acid, diisopropylsalicylic acid, octanoic acid, abietic acid, naphthenic acid, octanoic acid, lauric acid, tallic acid, and the like. Preferred  
10 positive charge directors are the metallic carboxylates (soaps) described in U.S. Patent 3,411,936 which include alkaline earth- and heavy-metallic salts of fatty acids containing at least 6-7 carbons and cyclic aliphatic acids including naphthenic acid; more preferred are polyvalent metal soaps of zirconium and aluminum; most preferred is the zirconium soap of octanoic acid (Zirconium HEX-CEM from Mooney  
15 Chemicals, Cleveland, OH).

The preferred charge direction levels for a given toner formulation will depend upon a number of factors, including the composition of the graft stabilizer and organosol, the molecular weight of the organosol, the particle size of the organosol, the core/shell ratio of the organosol, the pigment used in making the toner, and the  
20 ratio of organosol to pigment. In addition, preferred charge direction levels will also depend upon the nature of the electrophotographic imaging process, particularly the design of the developing hardware and photoreceptive element. Those skilled in the art, however, know how to adjust the level of charge direction based on the listed parameters to achieve the desired results for their particular application.

25 The conductivity of a liquid toner has been well established in the art as a measure of the effectiveness of a toner in developing electrophotographic images. A range of values from  $1.0 \times 10^{-11}$  mho/cm to  $10.0 \times 10^{-11}$  mho/cm has been disclosed as advantageous in U.S. 3,890,240. High conductivities generally indicate inefficient association of the charges on the toner particles and are seen in the low relationship  
30 between current density and toner deposited during development. Low conductivities indicate little or no charging of the toner particles and lead to very low development rates. The use of charge director compounds to ensure sufficient charge associated with

each particle is a common practice. There has, in recent times, been a realization that even with the use of charge directors there can be much unwanted charge situated on charged species in solution in the carrier liquid. Such charge produces inefficiency, instability, and inconsistency in the development. U.S. Patent No. 4,925,766 discloses  
5 that at least 40%, and preferably at least 80%, of the total charge in the liquid toner should be situated and remain on the toner particles.

Suitable efforts to localize the charges onto the toner particles and to ensure that there is substantially no migration of charge from those particles into the liquid, and that no other unwanted charge moieties are present in the liquid, give substantial  
10 improvements. A measure of the required properties is the ratio between the conductivity of the carrier liquid as it appears in the liquid toner and the conductivity of the liquid toner as a whole. This ratio is preferably less than 0.6, more preferably less than 0.4, and most preferably less than 0.3.

Any number of methods may be used for effecting particle size reduction of  
15 the pigment in preparation of the gel liquid toners. Some suitable methods include high shear homogenization, ball-milling, attritor milling, high energy bead (sand) milling, basket milling, or other means known in the art. Preferably, the ink is milled in an attritor, vertical bead mill, or basket mill to avoid overshearing the organosol, which can cause an undesirable reduction in ink stability and adversely affect charge  
20 characteristics such as ink conductivity.

In electrophotographic and electrographic processes, an electrostatic image is formed on the surface of a photoreceptive element or dielectric element. The photoreceptive element or dielectric element may be an intermediate transfer drum or belt, or the substrate for the final toned image itself, as described by Schmidt, S. P.  
25 and Larson, J. R. in Handbook of Imaging Materials Diamond, A. S., Ed: Marcel Dekker: New York; Chapter 6, pp 227-252, and U. S. Patent Nos. 4,728,983, 4,321,404, and 4,268,598.

In electrography, a latent image is typically formed by (1) placing a charge image onto the dielectric element (typically the receiving substrate) in selected areas  
30 of the element with an electrostatic writing stylus or its equivalent to form a charge image, (2) applying toner to the charge image, and (3) fixing the toned image. An example of this type of process is described in U.S. Patent No. 5,262,259.

Images formed by the present invention may be of a single color or a plurality of colors. Multicolor images can be prepared by repetition of the charging and toner application steps. Examples of electrophotographic methods suitable for producing full color reproductions are described by U.S. Patent Nos. 2,297,691; 2,752,833; 2,986,466; 3,690,756; 4,403,848; 4,370,047; 4,467,334; 4,728,983; U.S. 5,650,253; 5 2,916,718; and European Patent Application No. 0,453,256. Examples of suitable transfer and fixing processes are described in U.S. Patent Nos. 4,337,303 and 5,108,865.

In electrophotography, the electrostatic image is typically formed on a drum or 10 belt coated with a photoreceptive element by (1) uniformly charging the photoreceptive element with an applied voltage, (2) exposing and discharging portions of the photoreceptive element with a radiation source to form a latent image, (3) applying a toner to the latent image to form a toned image, and (4) transferring the toned image through one or more steps to a final receptor sheet. In some applications, 15 it is sometimes desirable to fix the toned image using a heated pressure roller or other fixing methods known in the art.

While the electrostatic charge of either the toner particles or photoreceptive element may be either positive or negative, electrophotography is preferably carried out by dissipating charge on a positively charged photoreceptive element. Toner is 20 then applied to the regions in which the positive charge was dissipated using a liquid toner immersion development technique. This development may be accomplished by using a uniform electric field produced by a development electrode spaced near the photoreceptive element surface. A bias voltage is applied to the electrode intermediate to the initially charged surface voltage and the exposed surface voltage 25 level. The voltage is adjusted to obtain the required maximum density level and tone reproduction scale for halftone dots without any background deposited. Liquid toner is then caused to flow between the electrode and the photoreceptive element. The charged toner particles are mobile in the field and are attracted to the discharged areas on the photoreceptive element while being repelled from the undischarged non-image 30 areas. Excess liquid toner remaining on the photoreceptive element is removed by techniques well known in the art. Thereafter, the photoreceptive element surface may be force dried or allowed to dry at the ambient conditions.

Particularly useful electrophotographic processes for forming a multi-colored image on a receptor are described in U.S. 5,061,583, U.S. 5,650,253 and U.S. 5,916,718. The process disclosed in U.S. 5,650,253 basically involves the steps of (i) applying a uniform positive charge of approximately 700 volts on the surface of a photoreceptive element, (ii) exposing and partially discharging the surface of the photoreceptive element with a laser scanning device in an image-wise pattern to create a latent image, (iii) applying a liquid color toner to the latent image to form both a toned image and a uniform surface charge on the photoreceptive element, (iv) removing excess liquid toner, (v) drying the toned image, and (vi) transferring the toned image either directly or indirectly onto a final receptor. To form multi-colored images, steps (i) through (iv) are repeated until all the desired colors are formed on the photoreceptive element prior to transferring the images either directly or indirectly onto a final receptor.

The process disclosed in U.S. 5,916,718 is similar to that of U.S. U.S. 5,650,253, except that processes (ii) through (v) are repeated in forming a multi-colored image. Unlike conventional electrophotographic processes, this process of forming multi-colored images may be accomplished without erasing the residual charge and recharging the surface of the photoreceptive element prior to scanning and developing a subsequent image. The liquid toner of the present invention provides sufficient charge in the imaged areas to allow the creation of a subsequent latent image without erasing and recharging the surface. Alternatively, the process of U.S. 5,916,718 may be carried out in a manner such that steps (ii)-(vi) are repeated in forming a multi-colored image, in which case the multi-colored image may be built up on an intermediate transfer element, or may be built up on the final image receptor.

These methods make use of an offset transfer process incorporating as an element an intermediate transfer roller which is coated with a silicone or fluorosilicone elastomer composition, which is heated to between 80-100°C, and which applies a force of approximately 40-80 lb<sub>f</sub> across the entire contact zone with the photoreceptor. One suitable coating composition for the transfer roller is the Dow Corning 94-003 fluorosilicone elastomer heated to between 85-95°C. Preferred elastomeric coating compositions are disclosed in U.S. 5,965,314.

The substrate for receiving the image from either the photoreceptive element in electrophotographic printing or the dielectric element in electrostatic printing can be any commonly used receptor material, such as paper, coated paper, polymeric films and primed or coated polymeric films, particularly adhesive coated polymeric  
5 films. Suitable polymeric films include polyesters, plasticized and compounded polyvinyl chloride (PVC), acrylics, polyurethanes, polyethylene/acrylic acid copolymers, (e.g., commercially available under the trade designation SURLYN and polyvinyl butyrals. Commercially available composite materials such as those having the trade designations SCOTCHCAL, SCOTCHLITE, and PANAFLEX are also  
10 suitable for preparing substrates.

The transfer of the formed image from the surface of photoreceptor or dielectric element to the final receptor or transfer medium may be enhanced by the incorporation of a release-promoting material within the dispersed particles used to form the image. The incorporation of a silicone-containing material or a fluorine-  
15 containing material in the outer (shell) layer of the particle facilitates the efficient transfer of the image.

In multicolor electrographic imaging, the toners may be applied to the surface of the dielectric element or photoreceptive element in any order, but for colorimetric reasons, bearing in mind the inversion that occurs on transfer, it is sometimes  
20 preferred to apply the images in a specified order depending upon the transparency and intensity of the colors. A preferred order for direct imaging or double transfer process is yellow, magenta, cyan, and black; for a single transfer process, the preferred order is black, cyan, magenta, and yellow. Yellow is generally imaged first to avoid contamination from other toners and black is generally imaged last due to the  
25 black toner acting as a filter of the radiation source.

Overcoating of the transferred image may optionally be carried out to protect the image from physical damage and/or actinic damage. Compositions for overcoatings are well-known in the art and typically comprise a clear film-forming polymer dissolved or suspended in a volatile solvent. An ultraviolet light absorbing  
30 agent may optionally be added to the coating composition. Lamination of protective layers to the image-bearing surface is also well known in the art and may be used with this invention.

In order to function most effectively, liquid toners preferably have conductance values in the range of 50 to 1200 picomho-cm<sup>-1</sup> at their working concentrations. Liquid toners prepared according to the present invention preferably have conductance values of from 50 to 500 picomho-cm<sup>-1</sup> for a dispersion containing 3% by weight solids. Toner  
5 are usually prepared in a concentrated form to conserve storage space and reduce transportation costs. In order to use the toners in the printer, the concentrate is diluted with additional carrier liquid to give what is termed the working strength liquid toner.

The acid-base gel organosol provides an efficient method for formulating a high solids ink. The acid-base gel organosol will phase separate into two phases, one  
10 phase comprising primarily a portion of the carrier liquid and the other phase a concentrated gel organosol dispersion. Once separated, the gel organosol can be simply redispersed with mixing or by removal of the separated carrier liquid form a concentrate of the dispersion. A high solids ink for printing or replenishment may be  
15 produced by allowing the organosol to gel and then decanting or siphoning off the supernatant liquid, thus forming a concentrate of the gel polymer in the dispersant liquid. A variety of alternative methods for concentrating the organosol or ink are well known in the art, such as gravity settling, centrifugation, filtration, controlled flocculation, etc. The carrier liquid may be removed either prior to or after the  
20 addition of the pigment and/or charge director and either before or after milling the toner.

These and other aspects of the present invention are demonstrated in the illustrative examples that follow.

### EXAMPLES

#### Glossary of Chemical Abbreviations & Chemical Sources

The following raw materials were used to prepare the polymers in the examples that follow:

The catalysts used in the examples are Azobisisobutyronitrile (designated as AIBN, VAZO-64 available from DuPont Chemicals, Wilmington, DE); and Dibutyl  
30 Tin Dilaurate (designated as DBTDL, available from Aldrich Chemical Co.,

Milwaukee, WI). The monomers are all available from Scientific Polymer Products, Inc., Ontario, NY unless designated otherwise.

The monomers used in the examples are designated by the following abbreviations: Dimethylaminoethyl methacrylate (DMAEMA); Dimethyl-m-  
5 isopropenyl benzylisocyanate (TMI, available from CYTEC Industries, West Paterson, NJ); Ethyl Acrylate (EA); 2-Ethylhexyl Methacrylate (EHMA); 2-Hydroxyethyl Methacrylate (HEMA); Lauryl Methacrylate or Dodecyl Methacrylate (LMA); Methacrylic acid (MAA); Methyl Methacrylate (MMA); and Octadecyl Acrylate or Stearyl Acrylate (ODA).

10

### Analytical Test Methods

The following test methods were used to characterize the polymers and inks in the examples which follow:

#### Percent Solids of Graft Stabilizer, Organosol, and Liquid Toner

Percent solids of the graft stabilizer solutions, and the organosol and ink  
15 dispersions, were determined gravimetrically using a halogen lamp drying oven attachment to a precision analytical balance (Mettler Instruments Inc., Hightstown, NJ). Approximately two grams of sample were used in each determination of percent solids using this sample drydown method.

#### Graft Stabilizer Molecular Weight

20 Various properties of the graft stabilizer have been determined to be important to the performance of the stabilizer, including molecular weight and molecular weight polydispersity. Graft stabilizer molecular weight is normally expressed in terms of the weight average molecular weight ( $M_w$ ), while molecular weight polydispersity is given by the ratio of the weight average molecular weight to the number average  
25 molecular weight ( $M_w/M_n$ ). Molecular weight parameters were determined for graft stabilizers with gel permeation chromatography (GPC) using tetrahydrofuran as the carrier solvent. Absolute  $M_w$  was determined using a Dawn DSP-F light scattering detector (Wyatt Technology Corp, Santa Barbara, CA), while polydispersity was  
30 evaluated by ratioing the measured  $M_w$  to a value of  $M_n$  determined with an Optilab 903 differential refractometer detector (Wyatt Technology Corp, Santa Barbara, CA).



### Organosol Particle Size

Organosol particle size was determined by dynamic light scattering on a diluted toner sample (typically < 0.0001 g/ml) using a Malvern Zetasizer III Photon Correlation Spectrometer (Malvern Instruments Inc, Southborough, MA). The dilute  
5 samples were ultrasonicated for one minute at 100 watts and 20 kHz (kHz) prior to measurement. Dynamic light scattering provides a fast method of determining the particle translational diffusion coefficient, which can be related to the z-average particle diameter without detailed knowledge of the optical and physical properties (i.e. refractive index, density and viscosity) of the organosol. Details of the method  
10 are described in Chu (Chu, B., Laser Scattering Academic Press, NY 1974, 11A). Since the organosols are comprised of nearly monodisperse, uniform spherical particles, dynamic light scattering provides an absolute measure of particle size for particles having diameters between 25-2500 nm.

### Liquid Toner Properties

15 The characterization of a liquid toner requires the measurement of a number of physical and chemical properties of the toner, as well as direct evaluation of image quality obtained by developing the toner in a LEP imaging mechanism. The measured toner characteristics can be roughly broken down into size-related properties (particle size), charge-related properties (bulk and free phase conductivity,  
20 dynamic mobility and zeta potential, and charge/developed reflectance optical density (Q/ROD), a parameter which is directly proportional to the toner charge/mass.

### Particle Size

Toner particle size distributions were determined using a Horiba LA-900 laser diffraction particle size analyzer (Horiba Instruments, Inc, Irvine, CA). Toner  
25 samples were diluted approximately 1/500 by volume and sonicated for one minute at 150 watts and 20 kHz prior to measurement. Toner particle size was expressed on a number-average basis in order to provide an indication of the fundamental (primary) particle size of the ink particles.

### Toner Conductivity

30 The liquid toner conductivity (bulk conductivity,  $k_b$ ) was determined at approximately 18 Hz using a Scientifica model 627 conductivity meter (Scientifica

- Instruments, Inc., Princeton, NJ). In addition, the free (dispersant) phase conductivity ( $k_f$ ) in the absence of toner particles was also determined. Toner particles were removed from the liquid milieu by centrifugation at 5°C for 1-2 hours at 6,000 rpm (6,110 relative centrifugal force) in a Jouan MR1822 centrifuge (Winchester, VA).
- 5 The supernatant liquid was then carefully decanted, and the conductivity of this liquid was measured using a Scientifica Model 627 conductance meter. The percentage of free phase conductivity relative to the bulk toner conductivity was then determined as: 100% ( $k_f/k_b$ ).

#### Particle Mobility

- 10 Toner particle electrophoretic mobility (dynamic mobility) was measured using a Matec MBS-8000 Electrokinetic Sonic Amplitude Analyzer (Matec Applied Sciences, Inc., Hopkinton, MA). Unlike electrokinetic measurements based upon microelectrophoresis, the MBS-8000 instrument has the advantage of requiring no dilution of the toner sample in order to obtain the mobility value. Thus, it was
- 15 possible to measure toner particle dynamic mobility at solids concentrations actually preferred in printing. The MBS-8000 measures the response of charged particles to high frequency (1.2 MHz) alternating (AC) electric fields. In a high frequency AC electric field, the relative motion between charged toner particles and the surrounding dispersion medium (including counter-ions) generates an ultrasonic wave at the same
- 20 frequency of the applied electric field. The amplitude of this ultrasonic wave at 1.2 MHz can be measured using a piezoelectric quartz transducer; this electrokinetic sonic amplitude (ESA) is directly proportional to the low field AC electrophoretic mobility of the particles. The particle zeta potential can then be computed by the instrument from the measured dynamic mobility and the known toner particle size,
- 25 dispersant liquid viscosity, and liquid dielectric constant.

#### Particle Charge

- Toner charge/mass is an important, albeit difficult to determine parameter useful in predicting the development characteristics (e.g. optical density, overtoning uniformity) for liquid toners. The difficulty in determining charge/mass for liquid
- 30 toners arises from the low developed toner mass (typically 50-200 micro grams/cm<sup>2</sup>) associated with the desired developed optical densities (typically > 1.2 reflectance optical density units). A related parameter which is directly proportional to toner

charge/mass is the toner charge/developed optical density. This parameter was determined by plating ink particles in distinct bands covering a range of known plating potentials onto a dielectric sheet coated with a silicone release layer while simultaneously monitoring the total current flow with a sensitive electrometer. The  
5 resulting plated toner layer was then air dried and transferred using an offset transfer process to plain paper. The reflectance optical density of the completely transferred toner film on paper was determined using a Greteq SPM50 reflectance optical densitometer (Greteq Instruments Inc., Regensdorf, Sitzerland). The ratio of the total current to the product of the plated toner area and the developed optical density yields  
10 the charge/ROD value for that toner, i.e.  $\text{Charge/ROD} = (\text{Total Current})/[(\text{Plated Area})(\text{Reflectance Optical Density})]$ .

### Graft Stabilizers

In the following examples of graft stabilizer preparations, it will be convenient to summarize the compositional details of each particular graft stabilizer or graft  
15 stabilizer precursor by ratioing the weight percentage of monomers employed in the synthesis. For example, a graft stabilizer designated LMA/HEMA-TMI (96/1/3-4.7 %w/w) is made from a graft stabilizer precursor that is a copolymer consisting of 96% weight percent LMA and 3% weight percent HEMA, to which is covalently bonded a grafting site consisting of 4.7 weight percent TMI based on the total weight of the  
20 graft stabilizer precursor.

### Example 1

A 3000 ml 3-necked round bottom flask equipped with a condenser, a thermocouple connected to a digital temperature controller, a nitrogen inlet tube  
25 connected to a source of dry nitrogen, and an overhead mechanical agitator, was charged with a mixture of 1468 g of NORPAR 12, 485.0 g of LMA, 16.0 g of 96% HEMA, and 5.0 g of AIBN. While mechanically stirring the mixture, the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/min. A hollow glass stopper was then inserted into the open end of the  
30 condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/min. The mixture was heated to 70°C with stirring, and the mixture was allowed to polymerize at 70°C for 16 hours, at which time the conversion was quantitative.

The mixture was heated to 90°C and held at that temperature for 1 hour to destroy any residual AIBN, then was cooled back to 70°C. The nitrogen inlet tube was then removed, and 7.8 g of 95% DBTDL were added to the mixture, followed by 23.5 g of TMI. The TMI was added drop wise over the course of approximately five minutes while magnetically stirring the reaction mixture. The nitrogen inlet tube was replaced, the hollow glass stopper in the condenser was removed, and the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/min. The hollow glass stopper was reinserted into the open end of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/min. The mixture was allowed to react at 70°C for 6 hours, at which time the conversion was quantitative.

The mixture was then cooled to room temperature. The cooled mixture was a viscous, colorless liquid containing no visible insoluble material. The percent solids of the liquid mixture was determined as 24.7% using the infrared drying method described above. Subsequent determination of molecular weight was made using GPC method described above. The copolymer had a  $M_w$  of 174,504 Da and a  $M_w/M_n$  of 4.24. The product is a copolymer of LMA and HEMA having random side chains of TMI and is designated herein as LMA/HEMA-TMI (97/3-4.7% w/w).

#### Example 2

Using the method and apparatus of Example 1, 1448 g of NORPAR 12, 485.0 g of LMA, 16.0 g of 96% HEMA, 15.0 g of MAA, and 5.0 g of AIBN were combined in a reaction flask. The resulting mixture reacted at 70°C for 16 hours. The mixture was heated to 90°C for 1 hour to destroy any residual AIBN, then was cooled back to 70°C. To the cooled mixture was then added 7.8 g of 95% DBTDL and 23.5 g of TMI. The TMI was added drop wise over the course of approximately five minutes while magnetically stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70°C for approximately 6 hours, at which time the conversion was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, colorless liquid containing no visible insoluble material.

The percent solids of the liquid mixture was determined as 26.4% using the infrared drying method described above. Subsequent determination of molecular weight was made using GPC method described above. The copolymer had a  $M_w$  of

213,719 Da and a  $M_w/M_n$  of 5.51. The product is a terpolymer containing LMA, MAA, and HEMA having random side chains of TMI and is designated herein as LMA/MAA/HEMA-TMI (97/3/3-4.7% w/w).

### Example 3

5        Using the method and apparatus of Example 1, 1448 g of NORPAR 12, 485.0 g of LMA, 16.0 g of 96% HEMA, 15.0 g of DMAEMA, and 5.0 g of AIBN were combined in a reaction flask. The resulting mixture reacted at 70°C for 16 hours. The mixture was heated to 90°C for 1 hour to destroy any residual AIBN, then was cooled back to 70°C. To the cooled mixture was then added 7.8 g of 95% DBTDL and 23.5  
10    g of TMI. The TMI was added drop wise over the course of approximately five minutes while magnetically stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70°C for approximately 6 hours, at which time the conversion was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, colorless liquid containing no visible insoluble  
15    material.

      The percent solids of the liquid mixture was determined as 25.6% using the infrared drying method described above. Subsequent determination of molecular weight was made using GPC method described above. The copolymer had a  $M_w$  of 185,331 Da and a  $M_w/M_n$  of 5.40. The product is a terpolymer containing LMA,  
20    DMAEMA, and HEMA having random side chains of TMI and is designated herein as LMA/DMAEMA/HEMA-TMI (94/3/3-4.7% w/w).

### Example 4

      A 16 ounce (0.48 liter), narrow-mouthed glass bottle was charged with 93.2 g of NORPAR 12, 30.6 g of ODA, 0.98 g of 96% HEMA, and 0.23 g of AIBN. The  
25    bottle was purged for two minutes with dry nitrogen at a rate of 1.5 liters/min, then sealed with a screw cap fitted with a teflon liner. The cap was secured in place using electrical tape. The sealed bottle was then inserted into a metal cage assembly and installed on the agitator assembly of an Atlas Launder-Ometer (Atlas Electrical  
Devices Company, Chicago, IL). The Launder-Ometer was operated at its fixed  
30    agitation speed of 42 rpm with a water bath temperature of 70°C. The mixture was allowed to react for 16 hours, at which time the conversion of monomer to polymer

was quantitative. The mixture was heated to 90°C for 1 hour to destroy any residual AIBN, then was cooled to room temperature. The bottle was then opened and 0.78 g of 95% DBTDL and 2.35 g of TMI were added to the cooled mixture. The bottle was purged for two minutes with dry nitrogen at a rate of approximately 1.5 liters/minute, then sealed with a screw cap fitted with a teflon liner. The cap was secured in place using electrical tape. The sealed bottle was then inserted into a metal cage assembly and installed on the agitator assembly of the Atlas Launder-Ometer. The Lander-Ometer was operated at its fixed agitation speed of 42 rpm with a water bath temperature of 70°C. The mixture was allowed to react for 6 hours, at which time the conversion was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, transparent straw colored solution that contained no insoluble solid residues and solidified slowly at room temperature.

The percent solids of the liquid mixture was determined as 26.1% using the infrared drying method described above. Subsequent determination of molecular weight was made using the GPC method described above. The copolymer had a  $M_w$  of 152,233 Da and a  $M_w/M_n$  of 6.19. The product is a copolymer of ODA and HEMA having random side chains of TMI and is designated herein as ODA/HEMA-TMI (94/3-4.7% w/w).

20

#### Example 5

Using the method and apparatus of Example 4, 93.2 g of NORPAR 12, 29.7 g of ODA, 0.95 g of DMAEMA, 0.98 g of 96% HEMA, and 0.23 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours. The mixture was then heated for to 90°C for 1 hour to destroy any residual AIBN, then was cooled to room temperature. To the cooled mixture was then added 0.78 g of 95% DBTDL and 2.35 g of TMI, and the mixture heated to 70°C and maintained at this temperature for 6 hours, after which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, straw colored solution that contained no insoluble solid residues and solidified slowly at room temperature.

30

The percent solids of the liquid mixture was determined as 26.2% using the infrared drying method described above. Subsequent determination of molecular

weight was made using the GPC method described above. The copolymer had a  $M_w$  of 106,690 Da and a  $M_w/M_n$  of 5.19. The product is a terpolymer of ODA, DMAEMA, and HEMA having random side chains of TMI and is designated herein as ODA/DMAEMA/HEMA-TMI (94/3/3-4.7% w/w).

5

#### Example 6

Using the method and apparatus of Example 4, 93.2 g of NORPAR 12, 29.7 g of ODA, 0.94 g of MAA, 0.98 g of 96% HEMA, and 0.23 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours. The mixture was then heated for to 90°C for 1 hour to destroy any residual AIBN, then  
10 was cooled to room temperature. To the cooled mixture was then added 0.78 g of 95% DBTDL and 2.35 g of TMI, and the mixture heated to 70°C and maintained at this temperature for 6 hours, after which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, slightly cloudy colorless solution that contained no insoluble solid residues and  
15 solidified slowly at room temperature.

The percent solids of the liquid mixture was determined as 25.7% using the infrared drying method described above. Subsequent determination of molecular weight was not possible due to unsuccessful filtration of the solution. The product is a terpolymer of ODA, MAA, and HEMA having random side chains of TMI and is  
20 designated herein as ODA/MAA/HEMA-TMI (94/3/3-4.7% w/w).

#### Example 7

A 5000 ml 3-necked round bottom flask equipped with a condenser, a thermocouple connected to a digital temperature controller, a nitrogen inlet tube connected to a source of dry nitrogen, and an overhead mechanical agitator, was  
25 charged with a mixture of 2563 g of NORPAR 12, 849 g of EHMA, 27 g of 96% HEMA, and 6.6 g of AIBN. While mechanically stirring the mixture, the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/min. A hollow glass stopper was then inserted into the open end of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/min.  
30 The mixture was heated to 70°C with stirring, and the mixture was allowed to polymerize at 70°C for 16 hours, at which time the conversion was quantitative.

The mixture was heated to 90°C and held at that temperature for 1 hour to destroy any residual AIBN, then was cooled back to 70°C. The nitrogen inlet tube was then removed, and 13.6 g of 95% DBTDL were added to the mixture, followed by 41.1 g of TMI. The TMI was added drop wise over the course of approximately  
5 five minutes while magnetically stirring the reaction mixture. The nitrogen inlet tube was replaced, the hollow glass stopper in the condenser was removed, and the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/min. The hollow glass stopper was reinserted into the open end  
10 of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/min. The mixture was allowed to react at 70°C for 6 hours, at which time the conversion was quantitative.

The mixture was then cooled to room temperature. The cooled mixture was a viscous, colorless liquid containing no visible insoluble material. The percent solids of the liquid mixture was determined as 25.5% using the infrared drying method  
15 described above. Subsequent determination of molecular weight was made using GPC method described above. The copolymer had a  $M_w$  of 220,197 Da and a  $M_w/M_n$  of 3.72. The product is a copolymer of EHMA and HEMA having random side chains of TMI and is designated herein as EHMA/HEMA-TMI (97/3-4.7% w/w).

20

### Example 8

Using the method and apparatus of Example 4, 74.5 g of NORPAR 13, 24.5 g of LMA, 0.78 g of 96% HEMA, and 0.188 g of AIBN were combined in a glass bottle, purged with nitrogen, and the resulting mixture heated at 70°C for 16 hours. The mixture was then heated for to 90°C for 1 hour to destroy any residual AIBN,  
25 then was cooled to room temperature. To the cooled mixture was then added 0.78 g of 95% DBTDL and 2.35 g of TMI, and the mixture heated to 70°C and maintained at this temperature for 6 hours, after which time the reaction was quantitative. The cooled mixture was a viscous, colorless solution that contained no insoluble solid residues and solidified slowly at room temperature.

30

The percent solids of the liquid mixture was determined as 25.1% using the infrared drying method described above. Subsequent determination of molecular weight was made using GPC method described above. The copolymer had a  $M_w$  of 211,918 Da and a  $M_w/M_n$  of 4.21. The product is a copolymer of LMA and HEMA



having random side chains of TMI and is designated herein as LMA/HEMA-TMI (97/3-4.7% w/w).

#### Example 9

Using the method and apparatus of Example 4, 74.5 g of NORPAR 13, 23.7 g  
5 of LMA, 0.76 g of MAA, 0.78 g of 96% HEMA, and 0.188 g of AIBN were  
combined in a glass bottle, purged with nitrogen, and the resulting mixture heated at  
70°C for 16 hours. The mixture was then heated for to 90°C for 1 hour to destroy any  
residual AIBN, then was cooled to room temperature. To the cooled mixture was then  
added 0.78 g of 95% DBTDL and 2.35 g of TMI, and the mixture heated to 70°C and  
10 maintained at this temperature for 6 hours, after which time the reaction was  
quantitative. The cooled mixture was a viscous, colorless solution that contained no  
insoluble solid residues and solidified slowly at room temperature.

The percent solids of the liquid mixture was determined as 25.8% using the  
infrared drying method described above. Subsequent determination of molecular  
15 weight was made using GPC method described above. The copolymer had a  $M_w$  of  
232,100 Da and a  $M_w/M_n$  of 3.82. The product is a terpolymer of LMA, MAA, and  
HEMA having random side chains of TMI and is designated herein as  
LMA/MAA/HEMA-TMI (94/3/3-4.7% w/w).

#### Organosol Examples

20 In the following examples of organosol preparation, it will be convenient to  
summarize the composition of each particular organosol in terms of the ratio of the  
total weight of monomers comprising the organosol core relative to the total weight of  
monomers comprising the organosol shell. The ratio is referred to as the core/shell  
25 ratio of the organosol. In addition it will be useful to summarize the compositional  
details of each particular organosol by ratioing the weight percentages of monomers  
used to create the shell and the core. For example, an organosol designated  
LMA/DMAEMA/HEMA-TMI/MMA/EA (94/3/3-4.7//25/75 % w/w) is made from  
shell comprised of graft stabilizer precursor which is a terpolymer consisting of 94  
30 weight percent of LMA, 3 weight percent of DMAEMA and 3 weight percent of  
HEMA, to which is covalently bonded a grafting site consisting of 4.7 weight percent  
of TMI bases on the total weight of graft stabilizer precursor. The graft stabilizer is

covalently bonded to a core that is comprised of 25 weight percent MMA and 75 weight percent of EA.

#### Example 10

This example uses the graft stabilizer of Example 1 to prepare a self stable organosol with a core/shell ratio of 8/1 having acidic groups (MAA, 1 weight percent w/w) in the core. An 8 ounce (0.24 liter), narrow-mouthed glass bottle was charged with 83.7 g of NORPAR 12, 8.0 g of EA, 2.7 g of MMA, 0.11 g of MAA (1 weight percent w/w), 5.4 g of the graft stabilizer mixture from Example 1 at 24.7%, solids, and 0.18 g of AIBN. The bottle was purged for two minutes with dry nitrogen at a rate of 1.5 liters/min, then sealed with a screw cap fitted with a teflon liner. The cap was secured in place using electrical tape. The sealed bottle was then inserted into a metal cage assembly and installed on the agitator assembly of an Atlas Launder-Ometer (Atlas Electrical Devices Company, Chicago, IL). The Launder-Ometer was operated at its fixed agitation speed of 42 rpm with a water bath temperature of 70°C. The mixture was heated to 70°C with agitation and allowed to polymerize at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion which did not gel.

This self stable organosol is designated LMA/HEMA-TMI//MAA/MMA/EA (97/3-4.7//1/25/74 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.5% using the halogen drying method outlined above. Subsequent determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 263 nm.

#### Example 11

This example uses the graft stabilizer of Example 1 to prepare a self stable organosol with a core/shell ratio of 8/1, having basic groups (DMAEMA, 3 weight percent w/w) in the core. A 3000 ml 3-necked round bottom flask equipped with an overhead mechanical stirrer, condenser, a thermocouple connected to a digital temperature controller, and a nitrogen inlet tube connected to a source of dry nitrogen, was charged with a mixture of 1511 g of NORPAR 12, 97.2 g of the graft stabilizer

from Example 1 at 24.7% solids, 47.0 g of MMA, 141.1 g of EA, 5.82 g of DMAEMA, and 3.24 g of AIBN. While mechanically stirring, the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/minute. The nitrogen flow rate was then adjusted to approximately 0.5  
5 liters/min. The mixture was heated to 70°C with stirring, and allowed to polymerize at 70°C for 16 hours at which time the organosol was cooled to room temperature.

Approximately 180 g of n-heptane were added to the cooled organosol and the resulting mixture was stripped of residual monomer using a rotary evaporator equipped with a dry ice/acetone condenser and operating at a temperature of 95°C,  
10 and reducing the vacuum gradually so as to maintain an adequate condensate collection rate to approximately 15 mm Hg. The stripped organosol was cooled to room temperature, yielding an opaque white dispersion that did not gel.

This self stable organosol is designated LMA/HEMA-TMI//DMAEMA/MMA/EA (97/3-4.7//3/24/73% w/w) having a calculated  $T_g$  equal  
15 to -1°C. The percent solids of this organosol was determined as 16.2% using the halogen drying method outlined above. Subsequent determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 140 nm.

### Example 12

20 This example use the graft stabilizer of Example 3 to prepare a self stable organosol with a core/shell ratio of 8/1, having basic groups (DMAEMA, 3 weight percent w/w) in the shell. Using the method and apparatus of Example 10, 83.9 g of NORPAR 12, 8.1 g of EA, 2.7 g of MMA, 5.2 g of the graft stabilizer solution from  
25 Example 3 at 25.6% solids, and 0.18 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated LMA/DMAEMA/HEMA-TMI//MMA/EA (94/3/3-4.7//25/75 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.8% using the halogen drying  
30 method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 163 nm.

### Example 13

- This example uses the graft stabilizer of Example 2 to prepare a self stable organosol with a core/shell ratio of 8/1, having acidic groups (MAA, 3 weight percent w/w) in the shell. Using the method and apparatus of Example 11, 1514 g of
- 5 NORPAR 12, 144.0 g of EA, 48.0 g of MMA, 90.9 g of the graft stabilizer solution from Example 2 at 26.4% solids, and 3.24 g of AIBN were reacted at 70°C for 16 hours. After addition of heptane and removal of the residual monomers under vacuum, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion that did not gel.
- 10 This self stable organosol is designated LMA/MAA/HEMA-TMI//MMA/EA (94/3/3-4.7//25/75 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 14.5% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 177
- 15 nm.

### Example 14

- This example uses the graft stabilizer of Example 4 to prepare a self stable organosol with a core/shell ratio of 8/1 having acidic groups (MAA, 1 weight percent w/w) in the core. A 16 ounce (0.48 liter), narrow-mouthed glass bottle was charged
- 20 with 168.1 g of NORPAR 12, 16.0 g of EA, 5.3 g of MMA, 0.22 g of MAA (1 weight percent w/w), 10.2 g of the graft stabilizer mixture from Example 4 at 26.1%, solids, and 0.36 g of AIBN. The bottle was purged for two minutes with dry nitrogen at a rate of 1.5 liters/min, then sealed with a screw cap fitted with a teflon liner. The cap was secured in place using electrical tape. The sealed bottle was then inserted into a
- 25 metal cage assembly and installed on the agitator assembly of an Atlas Launder-Ometer (Atlas Electrical Devices Company, Chicago, IL). The Launder-Ometer was operated at its fixed agitation speed of 42 rpm with a water bath temperature of 70°C. The mixture was heated to 70°C with agitation and allowed to polymerize at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an
- 30 opaque white dispersion that did not gel.

This self stable organosol is designated ODA/HEMA-TMI//MAA/MMA/EA (97/3-4.7//1/25/74 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of

this organosol was determined as 11.9% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 146 nm.

5

#### Example 15

This example uses the graft stabilizer of Example 4 to prepare a self stable organosol with a core/shell ratio of 8/1, having basic groups (DMAEMA, 3 weight percent w/w) in the core. Using the method and apparatus of Example 14, 168.4 g of NORPAR 12, 15.8 g of EA, 5.2 g of MMA, 0.65 g of DMAEMA (3 weight percent w/w), 10.2 g of the graft stabilizer solution from Example 4 at 26.1% solids, and 0.36 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated ODA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/1/3/24/73% w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.4% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 201 nm.

20

#### Example 16

This example uses the graft stabilizer of Example 5 to prepare a self stable organosol with a core/shell ratio of 8/1, having basic groups (DMAEMA, 3 weight percent w/w) in the shell. Using the method and apparatus of Example 14, 168.0 g of NORPAR 12, 16.1 g of EA, 5.4 g of MMA, 10.2 g of the graft stabilizer solution from Example 5 at 26.2% solids, and 0.36 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated ODA/DMAEMA/HEMA-TMI/MMA/EA (94/3/3-4.7/1/25/75 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.7% using the halogen drying method outlined above. Determination of average particle size was made using the

dynamic light scattering method described above; the organosol had a z-average diameter of 156 nm.

#### Example 17

This example uses the graft stabilizer of Example 5 to prepare a self stable organosol with a core/shell ratio of 8/1 having acidic groups (MAA, 1 weight percent w/w) and basic groups (DMAEMA, 3 weight percent w/w) in the core and shell, respectively. Using the method and apparatus of Example 14, 168.3 g of NORPAR 12, 16.0 g of EA, 5.3 g of MMA, 0.22 g of MAA (1 weight percent w/w), 10.2 g of the graft stabilizer solution from Example 5 at 26.2% solids, and 0.36 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated ODA/DMAEMA/HEMA-TMI/MAA/MMA/EA (94/3/3-4.7//1/25/74 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.7% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 248 nm.

#### Example 18

This example uses the graft stabilizer of Example 6 to prepare a self stable organosol with a core/shell ratio of 8/1, having acidic groups (MAA, 3 weight percent w/w) in the shell. Using the method and apparatus of Example 14, 167.8 g of NORPAR 12, 16.1 g of EA, 5.4 g of MMA, 10.4 g of the graft stabilizer solution from Example 6 at 25.7% solids, and 0.36 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated ODA/MAA/HEMA-TMI/MMA/EA (94/3/3-4.7//1/25/75 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.5% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light

scattering method described above; the organosol had a z-average diameter of 155 nm.

#### Example 19

This example uses the graft stabilizer of Example 8 to prepare a self stable organosol with a core/shell ratio of 8/1, having basic groups (DMAEMA, 3 weight percent w/w) in the core in a solvent having a Hildebrand solubility parameter lower than that of NORPAR 12. Using the method and apparatus of Example 10, 84.1 g of NORPAR 13, 7.8 g of EA, 2.6 g of MMA, 0.32 g of DMAEMA (3 weight percent w/w), 5.3 g of the graft stabilizer solution from Example 8 at 25.1% solids, and 0.18 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated LMA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/1/3/24/73 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 12.0% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 232 nm.

#### Example 20

This example uses the graft stabilizer of Example 9 to prepare a self stable organosol with a core/shell ratio of 8/1, having acidic groups (MAA, 3 weight percent w/w) in the shell in a solvent having a Hildebrand solubility parameter lower than that of NORPAR 12. Using the method and apparatus of Example 10, 84.1 g of NORPAR 13, 8.1 g of EA, 2.7 g of MMA, 5.1 g of the graft stabilizer solution from Example 9 at 25.8% solids, and 0.18 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that did not gel.

This self stable organosol is designated LMA/MAA/HEMA-TMI/MMA/EA (94/3/3-4.7/1/25/75 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.6% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light

scattering method described above; the organosol had a z-average diameter of 159 nm.

#### Example 21

This example uses the graft stabilizer of Example 2 to prepare a gel organosol with a core/shell ratio of 8/1 having both basic groups (DMAEMA, 3 weight percent w/w) and acidic groups (MAA, 3 weight percent w/w) in the core and shell, respectively. Using the method and apparatus of Example 10, 84.3 g of NORPAR 12, 7.9 g of EA, 2.6 g of MMA, 0.32 g of DMAEMA (3 weight percent w/w), 5.1 g of the graft stabilizer solution from Example 2 at 26.4% solids, and 0.18 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at 70°C for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that formed a moderately strong gel over a period of 24 hours.

This gel organosol is designated LMA/MAA/HEMA-TMI/DMAEMA/MMA/EA (94/3/3-4.7/1/3/24/73 % w/w) having a calculated  $T_g$  equal to -1°C. The percent solids of this organosol was determined as 11.8% using the halogen drying method outlined above. Accurate determination of average particle size was not possible using the dynamic light scattering method described owing to rapid flocculation of the particles during the time scale of the analysis.

#### Example 22

This example uses the graft stabilizer of Example 7 to prepare a gel organosol with a core/shell ratio of 8/1 having basic groups (DMAEMA, 3 weight percent w/w) in the core. A 16 ounce (0.48 liter), narrow-mouthed glass bottle was charged with 168.2 g of NORPAR 12, 15.8 g of EA, 5.23 g of MMA, 0.65 g of DMAEMA (3 weight percent w/w), 10.46 g of the graft stabilizer mixture from Example 7 at 25.5% solids, and 0.36 g of AIBN. The bottle was purged for two minutes with dry nitrogen at a rate of 1.5 liters/min, then sealed with a screw cap fitted with a teflon liner. The cap was secured in place using electrical tape. The sealed bottle was then inserted into a metal cage assembly and installed on the agitator assembly of an Atlas Launder-Ometer (Atlas Electrical Devices Company, Chicago, IL). The Launder-Ometer was operated at its fixed agitation speed of 42 rpm with a water bath temperature of 70°C. The mixture was heated to 70°C with agitation and allowed to polymerize at 70°C for 16 hours, at which time the organosol was cooled to room



temperature yielding an opaque white dispersion that formed a moderately strong gel over 4 hours.

This gel organosol is designated EHMA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/3/24/73 % w/w) having a calculated  $T_g$  equal to  $-1^\circ\text{C}$ . The percent solids of this organosol was determined as 11.9% using the halogen drying method outlined above. Determination of average particle size was made using the dynamic light scattering method described above; the organosol had a z-average diameter of 155 nm.

#### Example 23

This example uses the graft stabilizer of Example 7 to prepare a gel organosol with a core/shell ratio of 8/1 having acidic groups (MAA, 1 weight percent w/w) in the core. Using the method and apparatus of Example 14, 167.8 g of NORPAR 12, 16.0 g of EA, 5.33 g of MMA, 0.215 g of MAA (1 weight percent w/w), 10.46 g of the graft stabilizer mixture from Example 7 at 25.5% solids, and 0.36 g of AIBN were combined in a reaction bottle and the resulting mixture reacted at  $70^\circ\text{C}$  for 16 hours, at which time the organosol was cooled to room temperature yielding an opaque white dispersion that formed a moderately strong gel over a period of 2 hours.

This gel organosol is designated EHMA/HEMA-TMI/MAA/MMA/EA (97/3-4.7/1/25/74 % w/w) having a calculated  $T_g$  equal to  $-1^\circ\text{C}$ . The percent solids of this organosol was determined as 11.6% using the halogen drying method outlined above. Accurate determination of average particle size was not possible using the dynamic light scattering method described owing to settling of the particles during the time scale of the analysis.

#### Example 24

20 g of the self stable organosol from Example 10, LMA/HEMA-TMI/MAA/MMA/EA (97/3-4.7/1/25/74 % w/w), was mixed with 20 g of the self stable organosol from Example 11, LMA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/3/24/73% w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 25

20 g of the self stable organosol from Example 14, ODA/HEMA-TMI/MAA/MMA/EA (97/3-4.7/1/25/75 % w/w), was mixed with 20 g of the self stable organosol from Example 15, ODA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/3/24/73% w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 26

20 g of the self stable organosol from Example 12, LMA/DMAEMA/HEMA-TMI/MMA/EA (94/3/3-4.7/25/75 % w/w), was mixed with 20 g of the self stable organosol from Example 13, LMA/MAA/HEMA-TMI/MMA/EA (94/3/3-4.7/25/75 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 27

20 g of the self stable organosol from Example 16, ODA/DMAEMA/HEMA-TMI/MMA/EA (94/3/3-4.7/25/75 % w/w), was mixed with 20 g of the self stable organosol from Example 18, ODA/MAA/HEMA-TMI/MMA/EA (94/3/3-4.7/25/75 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 28

14.8 g of the self stable organosol from Example 11 at 16.2% solids, LMA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/3/24/73% w/w), was mixed with 16.6 g of the self stable organosol from Example 13 at 14.5% solids, LMA/MAA/HEMA-TMI/MMA/EA (94/3/3-4.7/25/75 % w/w), and 8.6 g of NORPAR 12 in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 29

20 g of the self stable organosol from Example 18, ODA/HEMA-TMI/DMAEMA/MMA/EA (97/3-4.7/3/24/73% w/w), was mixed with 20 g of the self stable organosol from Example 15, ODA/MAA/HEMA-TMI/MMA/EA (94/3/3-

4.7//25/75 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Example 30

- 20 g of the self stable organosol from Example 14, ODA/HEMA-TMI//MAA/MMA/EA (97/3-4.7//1/25/74% w/w), was mixed with 20 g of the self stable organosol from Example 16, ODA/DMAEMA/HEMA-TMI//MMA/EA (94/3/3-4.7//25/75 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a weak gel on standing over a period of 72 hours.

#### Example 31

- 20 g of the self stable organosol from Example 22, EHMA/HEMA-TMI//DMAEMA/MMA/EA (97/3-4.7//3/24/73% w/w), was mixed with 20 g of the self stable organosol from Example 23, EHMA/HEMA-TMI//MAA/MMA/EA (94/3/3-4.7//1/25/74 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a very strong gel over 1-2 hours.

#### Example 32

- 20 g of the self stable organosol from Example 19, LMA/HEMA-TMI//DMAEMA/MMA/EA (97/3-4.7//3/24/73% w/w), was mixed with 20 g of the self stable organosol from Example 20, LMA/MAA/HEMA-TMI//MMA/EA (94/3/3-4.7//25/75 % w/w), in a 4 ounce glass jar. The resultant organosol mixture formed a gel on standing over a period of 24 hours.

#### Liquid Toner Examples

- The following liquid toner examples were prepared with a gel organosol formed by mixing 741 g of LMA/HEMA-TMI//DMAEMA/MMA/EA (97/3-4.7//3/24/73% w/w) at 16.2% solids with 827 g of LMA/MAA/HEMA-TMI//MMA/EA (94/3/3-4.7//25/75 % w/w) at 14.5% solids, and 32 g of NORPAR 12 to produce the gel organosol mixture of Example 28 at 15.0% solids.

#### Example 33

This is an example of preparing a cyan liquid toner at an organosol/pigment ratio of 8 using the above-described gel organosol mixture. The organosol was mixed using a Silverson mixer (Model L22R, Silverson machines. Ltd, Waterside, England) operated at the lowest speed setting. After mixing for 5 minutes, 213.3 g of the  
5 homogenized organosol were combined with 88.9 g of NORPAR 12, 4 g of the Pigment Blue 15:3 (C.I 74160:3; #249-1282, Sun Chemical Company, Cincinnati, Ohio), and 3.90 g of 6.16% Zirconium HEX-CEM solution (OMG chemical company, Cleveland, Ohio) in an eight ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Aimex Co.,Ltd., Tokyo, Japan)  
10 charged with 390 g of 1.3mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2000rpm for 1.5 hours without cooling water circulating through the cooling jacket of the milling chamber.

A portion of this 12% (w/w) solids toner concentrate was diluted to approximately 3% (w/w). This dilute toner sample exhibited the following properties  
15 as determined using the test methods described above:

Number Mean Particle Size: 0.392 microns

Bulk Conductivity: 84 picoMhos/cm

Percent Free Phase Conductivity: 8%

Dynamic Mobility: 0.0277 micron-cm/[Volt-second]

20 Zeta Potential: 62 mV

This working strength toner was tested on the plating apparatus described previously. The reflection optical density (ROD) was greater than 0.9 at plating voltages greater than 450 volts.

#### Example 34

25 This is an example of preparing a yellow liquid toner at an organosol/pigment ratio of 6 using the above-described gel organosol mixture. The organosol was mixed using a Silverson mixer (Model L22R, Silverson machines. Ltd, Waterside, England) operated at the lowest speed setting. After mixing for 5 minutes, 205.7 g of the homogenized organosol in NORPAR 12 were combined with 86.6 g of NORPAR 12,  
30 5.14 g of Pigment Yellow 83 (C.I. 21108; #275-0570, Sun Chemical Company, Cincinnati, Ohio), and 2.5 g of 6.16% Zirconium HEX-CEM solution (OMG chemical company, Cleveland, Ohio) in an eight ounce glass jar. This mixture was

then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Aimex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2000rpm for 1.5 hours without cooling water circulating through the cooling jacket of the milling chamber.

5        A portion of this 12% (w/w) solids toner concentrate was diluted to approximately 3% (w/w). This dilute toner sample exhibited the following properties as determined using the test methods described above:

Number Mean Particle Size: 0.816 microns

Bulk Conductivity: 69 picoMhos/cm

10        Percent Free Phase Conductivity: 14%

Dynamic Mobility: 0.0304 micron-cm/[Volt-second]

Zeta Potential: 68 mV

      This working strength toner was tested on the plating apparatus described previously. The reflection optical density (ROD) was greater than 1.3 at plating  
15        voltages greater than 450 volts.

#### Example 35

      This is an example of preparing a black liquid toner at an organosol/pigment ratio of 6 using the above-described gel organosol mixture. The organosol was mixed using a Silverson mixer (Model L22R, Silverson machines. Ltd, Watreside, England)  
20        operated at the lowest speed setting. After mixing for 5 minutes, 205.7 g of the homogenized organosol in NORPAR 12 were combined with 85.8 g of NORPAR 12, 5.14 g of Monarch 120 carbon black (Cabot Corporation, Billerica, Mass.), and 3.34 g of 6.16% Zirconium HEX-CEM solution (OMG chemical company, Cleveland, Ohio) in an eight ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead  
25        mill (Model 6TSG-1/4, Aimex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2000rpm for 1.5 hours without cooling water circulating through the cooling jacket of the milling chamber.

      A portion of this 12% (w/w) solids toner concentrate was diluted to  
30        approximately 3% (w/w). This dilute toner sample exhibited the following properties as determined using the test methods described above:

Number Mean Particle Size: 0.288 microns

Bulk Conductivity: 96 picoMhos/cm

Percent Free Phase Conductivity: 4%

Dynamic Mobility: 0.0323 micron-cm/[Volt-second]

Zeta Potential: 72 mV

5 This working strength toner was tested on the plating apparatus described previously. The reflection optical density (ROD) was greater than 1.3 at plating voltages greater than 450 volts.

#### Example 36

10 This example illustrates the preparation of a high solids gel organosol from a preformed lower solids gel organosol. Such high solids organosols can be used to prepare high solids toner concentrates suitable for use as toner replenishments in liquid immersion development.

Approximately 1000 g of the gel organosol mixture of Example 28 at 15.0% solids was stored without agitation in a two liter wide mouth polyethylene bottle. After approximately 24 hours, the sample had separated into a lower gel phase and a  
15 clear supernatant liquid phase consisting essentially of NORPAR 12. The clear supernatant liquid was carefully decanted, and a sample of the lower gel phase was removed. The percent solids of this sample of concentrated gel organosol was determined using the infrared drying method described above. The gel organosol solids had increased from 15.0% to 35.1%. The resulting gel organosol concentrate  
20 can be used to prepare of gel inks at concentrations greater than 35% (w/w).

#### Example 37

This example illustrates the preparation of a high solids gel ink from a preformed lower solids gel ink. This example also illustrates a method for additionally reducing the free phase conductivity of such gel ink concentrates relative  
25 to the starting ink. Such high solids inks with reduced free phase conductivity are particularly useful as toner replenishers in liquid immersion development.

Approximately 35 g of the yellow gel ink at 12.0% solids (11% free phase conductivity) from Example 34 was transferred to a 50 ml centrifuge tube and centrifuged at 5°C. for 5 minutes at 7500rpm (7,640 relative centrifugal force) in a  
30 Jouan MR1822 centrifuge. The ink separated into a lower pigmented gel phase and a clear supernatant liquid phase consisting of essentially NORPAR 12. The clear

supernatant liquid was carefully decanted, and a sample of the lower gel phase was removed. The percent solids of this sample of concentrated gel toner was determined using the infrared drying method described above. The gel toner solids had increased from 12.0% to 28.1%. Approximately 10 g of NORPAR 12 was added to this gel  
5 toner concentrate. The gel concentrate was mixed by shaking the capped centrifuge tube. The concentrated gel ink readily redispersed upon mixing with the added NORPAR 12. This sample was centrifuged at 5°C. for 1 hour at 7500 rpm (7,640 relative centrifugal force) in the Jouan MR1822 centrifuge. The supernatant was then carefully decanted, and the conductivity of this solution was then measured using a  
10 Scientifica Model 627 conductivity meter operating at 12 Hz. The percentage of free phase conductivity relative to the bulk toner conductivity was now found to be approximately 1.0%.

A number of embodiments of the invention have been described.

15 Nevertheless, it will be understood that various modifications may be made without departing from the spirit and scope of the invention. Accordingly, other embodiments are within the scope of the following claims.

**WHAT IS CLAIMED IS:**

1. A gel organosol dispersion comprising:
  - (a) a carrier liquid having Kauri-Butanol number less than 30; and
  - (b) a graft copolymer comprising a (co)polymeric steric stabilizer
- 5 covalently bonded to a thermoplastic (co)polymeric core that is insoluble in said carrier liquid,  
wherein said graft copolymer comprises Lewis acid and Lewis base-functional groups that interact with each other to form non-covalent crosslinks.
2. A gel organosol dispersion according to claim 1 wherein said dispersion
- 10 further comprises a colorant.
3. A gel organosol dispersion according to any one of the preceding claims wherein said graft copolymer comprises Lewis acid-functional units selected from the group consisting acrylic acid, methacrylic acid, 2-acrylamido-2-methyl propane sulfonic acid, and combinations thereof.
- 15 4. A gel organosol dispersion according to any one of the preceding claims wherein said graft copolymer comprises 0.1-5% of said Lewis acid-functional units on a weight to weight basis.
5. A gel organosol dispersion according to any one of the preceding claims wherein said graft copolymer comprises Lewis base-functional units derived from
- 20 monomers selected from the group consisting of t-butylamino methacrylate, diethylaminoethyl methacrylate, 2-dimethylaminoethyl methacrylate, dimethylaminopropyl acrylamide, 2-diisopropylaminoethyl methacrylate, 4-hydroxybutyl acrylate, 2-hydroxyethyl acrylate, 2-hydroxyethyl methacrylate, 2-hydroxypropyl acrylate, 2-hydroxypropyl methacrylate, 2-vinyl pyridine, 4-vinyl
- 25 pyridine, N-vinyl pyridine, and combinations thereof.
6. A gel organosol dispersion according to any one of the preceding claims wherein said dispersion has a solids content of at least 2% by weight.



7. A gel organosol dispersion according to any one of the preceding claims wherein the Hildebrand solubility parameter difference between said stabilizer and said carrier liquid is no greater than  $3.0 \text{ mPa}^{1/2}$ .

8. A gel organosol dispersion according to claim 2 wherein said colorant  
5 comprises a pigment.

9. A gel organosol dispersion according to claim 8 wherein the ratio of said graft copolymer to said pigment on a weight to weight basis is between 1/1 and 20/1.

10. A gel organosol dispersion according to any one of the preceding claims wherein said dispersion further comprises a charge director.

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